

## CHEMISTRY

# Electronic structure of aqueous solutions: Bridging the gap between theory and experiments

Tuan Anh Pham,<sup>1\*</sup> Marco Govoni,<sup>2,3</sup> Robert Seidel,<sup>4†</sup> Stephen E. Bradforth,<sup>4</sup>  
Eric Schwegler,<sup>1</sup> Giulia Galli<sup>2,3</sup>

Predicting the electronic properties of aqueous liquids has been a long-standing challenge for quantum mechanical methods. However, it is a crucial step in understanding and predicting the key role played by aqueous solutions and electrolytes in a wide variety of emerging energy and environmental technologies, including battery and photoelectrochemical cell design. We propose an efficient and accurate approach to predict the electronic properties of aqueous solutions, on the basis of the combination of first-principles methods and experimental validation using state-of-the-art spectroscopic measurements. We present results of the photoelectron spectra of a broad range of solvated ions, showing that first-principles molecular dynamics simulations and electronic structure calculations using dielectric hybrid functionals provide a quantitative description of the electronic properties of the solvent and solutes, including excitation energies. The proposed computational framework is general and applicable to other liquids, thereby offering great promise in understanding and engineering solutions and liquid electrolytes for a variety of important energy technologies.

## INTRODUCTION

Liquid electrolytes are essential components of important devices used in a wide variety of emerging technologies, including solar water splitting, batteries, and supercapacitors, for energy conversion and storage (1–3). One of the essential prerequisites to predict and optimize the performance of these devices is the detailed understanding of the electronic properties of electrolytes, including their ionization potentials (IPs) and electron affinities (EAs). For example, a proper energy alignment at the electrode-electrolyte interface of photoelectrochemical (PEC) cells is key to achieving efficient hydrogen production (4–6). In addition, IPs and EAs of electrolytes determine their electrochemical windows and thereby the electrochemical stability of the electrode-electrolyte interface in lithium-ion batteries (7, 8). Establishing a predictive theory of the electronic properties of electrolytes, particularly aqueous solutions, that can be used to complement and guide experiments is therefore an important task.

Quantum mechanical methods that do not rely on a priori assumptions on the interatomic forces are excellently suited for this task. However, in practice, predicting the electronic structure of electrolytes based on quantum mechanics remains a great challenge because of the high computational complexity involved in the modeling of their finite temperature properties. In particular, first-principles molecular dynamics (MD) simulations of models consisting of at least hundreds of atoms are required to properly account for the structure of a liquid environment. In addition, an accurate and simultaneously efficient computational framework is desired for describing the electronic properties of the liquid.

Applications of highly accurate quantum chemistry methods, such as coupled-cluster theory, to liquid phases have not been possible because of the prohibitive computational cost and the difficulty of cou-

pling these methods with MD. On the other hand, semilocal density functional theory (DFT) widely used for the simulations of condensed phases often fails to provide a quantitative description of the electronic properties of electrolytes, including simple salt solutions (9). Alternative approaches included the integration of quantum chemistry methods with solvation models, such as polarizable continuum models (10–12) and the effective fragment potential method (13); the former treat the electronic properties of the species of interest in a quantum mechanical region that is embedded in a dielectric continuum described by solvation models. It has been shown that the inclusion of a large quantum mechanical region may be important to account for the properties of complex solutions (12), and the results may significantly depend on the choice of the solvation models (13).

At variance with liquid electrolytes, a systematic hierarchy of approximate yet highly successful first-principles techniques has been established to describe the electronic properties of both pristine and defective semiconductors and insulators in the solid state. Of particular interest are recent developments of hybrid exchange-correlation functionals that contain a fraction of exact exchange ( $\alpha$ ) proportional to the inverse of the electronic dielectric constant ( $\epsilon_\infty$ ) of the material [hereafter referred to as dielectric-dependent hybrid (DDH) functionals] (14–16). This class of functionals has been shown to provide a good description of the electronic properties not only of bulk materials but also of more complex systems, including band offsets of heterogeneous interfaces and defect charge transition levels in doped semiconductors (17–19). In addition, significant progress has been made in the past decade on post-DFT methods, including the development of efficient approaches for calculating the electronic properties of semiconductors in the framework of many-body perturbation theory (MBPT) (20). Owing to algorithmic advances (21, 22) and progress in high-performance computing, it has now become within reach to combine first-principles MD with hybrid functionals and MBPT to study liquid phases. Encouraging applications have been recently reported (23–27); however, the predictive power of these combined techniques has not been fully explored and systematically investigated.

Here, we present a predictive and efficient first-principles approach to describing the electronic properties of aqueous solutions. We extended first-principles techniques originally developed in solid-state

Copyright © 2017  
The Authors, some  
rights reserved;  
exclusive licensee  
American Association  
for the Advancement  
of Science. No claim to  
original U.S. Government  
Works. Distributed  
under a Creative  
Commons Attribution  
NonCommercial  
License 4.0 (CC BY-NC).

Downloaded from https://www.science.org at Universit degli Studi di Modena e Reggio Emilia on March 17, 2023

<sup>1</sup>Quantum Simulations Group, Lawrence Livermore National Laboratory, Livermore, CA 94550, USA. <sup>2</sup>Institute for Molecular Engineering, University of Chicago, Chicago, IL 60637, USA. <sup>3</sup>Materials Science Division, Argonne National Laboratory, Lemont, IL 60439, USA. <sup>4</sup>Department of Chemistry, University of Southern California, Los Angeles, CA 90089–0482, USA.

\*Corresponding author. Email: pham16@llnl.gov

†Present address: Methods for Material Development, Helmholtz-Zentrum Berlin für Materialien und Energie, D-12489 Berlin, Germany.

physics for the study of defects to investigate the photoelectron (PE) spectra of aqueous solutions, particularly of solvated anions, and we validated our theoretical results by direct comparison with state-of-the-art liquid-jet PE experiments. We show that a combination of first-principles MD and electronic structure calculations using DDH functionals yields an excellent description of the PE spectra for a broad range of aqueous solutions, providing electronic energy levels of liquid water and solvated anions in quantitative agreement with experiments. In addition, the DDH functionals show comparable performance to the more sophisticated but computationally more expensive MBPT methods. Finally, we demonstrate how these theoretical techniques can be coupled with liquid-jet experiments to provide a detailed interpretation of experimental PE spectra and a direct mapping of chemical species to specific features of the spectra. We note that the experimental technique used here only provides information on the occupied electronic energy levels of aqueous solutions. Unoccupied electronic states can be accessed using other techniques, such as near-edge x-ray absorption fine structure spectroscopy (NEXAFS), which probes the excitation of atomic core electrons to unoccupied molecular orbitals. Because these orbitals are sensitive to the local chemical environment, NEXAFS is more sensitive to the structure and bonding of solutions than x-ray PE spectroscopy, and it has been used to determine the solvation environments of several of the molecules studied here, such as  $\text{CO}_3^{2-}$ ,  $\text{HCO}_3^-$ , and  $\text{NO}_3^-$  (28, 29).

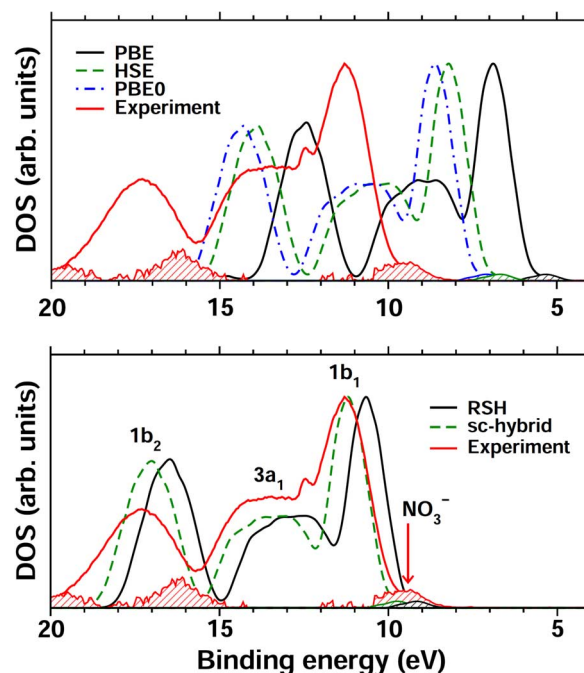
## RESULTS AND DISCUSSION

### Accuracy of the computational approach

We first address the performance of different electronic structure theories for a specific, representative ion in water, that is,  $\text{NO}_3^-$ , which is commonly used in aqueous electrolytes for high-voltage supercapacitors (30) and solar-to-fuel conversion (31). We focus on anions because their frontier electronic energy levels are usually either close to the band edges of liquid water or within its bandgap and hence directly influence the chemical reactivity of species in solutions. We note that liquid water can be considered as a wide bandgap insulator, and simple anions/cations in water play the same role as defects in insulators and semiconductors (32, 33).

In Fig. 1, we show the PE spectrum (red line) of a 1.0 M  $\text{NaNO}_3$  aqueous solution measured by liquid microjet techniques (see Materials and Methods), compared to the results obtained at different levels of theory. Several bands are present, corresponding to the binding energies (BEs) of valence electrons excited into vacuum. The bands are labeled (see bottom panel) according to the symmetry of single-particle states from which the emission occurs. Three peaks ( $1b_2$ ,  $3a_1$ , and  $1b_1$ ) are due to electrons ejected from water molecules; the peak corresponding to the  $\text{NO}_3^-$  highest occupied molecular orbital (HOMO) level is located at 9.42 eV (the full spectrum that includes the cation signal is reported in fig. S1).

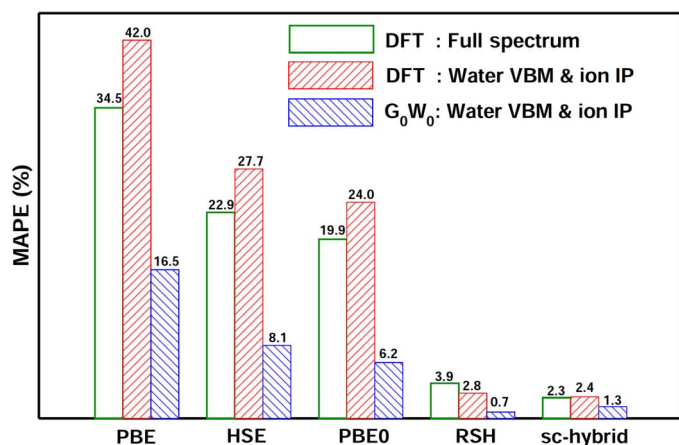
The calculated valence PE spectra shown in Fig. 1 were obtained as DFT Kohn-Sham energies with semilocal [Perdew, Burke, and Ernzerhof (PBE)] and hybrid [Heyd, Scuseria, and Ernzerhof (HSE) and PBE0] functionals (top panel); the latter are commonly used for examining the electronic properties of semiconductors (14, 34). In addition, we considered newly developed range-separated hybrid (RSH) and self-consistent hybrid (sc-hybrid) functionals (bottom panel), which have been shown to provide accurate results for excited-state properties of extended and finite systems (15, 16). The sc-hybrid and RSH functionals include a fraction of exact exchange equal to the



**Fig. 1. PE spectra of the  $\text{NO}_3^-$  solution.** The theoretical results were obtained with DFT using several functionals: PBE, HSE, and PBE0 (top panel), as well as RSH and sc-hybrid functionals (bottom panel). The experimental PE spectrum is indicated by the red solid line. The corresponding shaded areas represent the differential valence PE spectra obtained by subtracting the water spectrum from raw experimental  $\text{NaNO}_3$  spectra, and theoretical density of state (DOS) contributions from the  $\text{NO}_3^-$  HOMO. All energy levels are relative to vacuum. Experimental and theoretical DOS intensities were rescaled with respect to the  $1b_1$  peak of water at 11.31 eV. The differential spectra and theoretical HOMO DOS intensity were magnified (three times) for clarity.

inverse macroscopic dielectric constant of water, determined self-consistently from first-principles to be 0.6098. The attenuation parameter in the RSH functional was determined nonempirically and set to the Thomas-Fermi screening length of liquid water, that is,  $0.58 \text{ bohr}^{-1}$  (15, 16). We note that calculations of semiconductor bandgaps and vertical IPs of gas-phase molecules have shown a weak dependence of the results on the choice of the attenuation parameter entering the RSH functional (16).

As expected, the PBE functional severely underestimates experimental results, and the HSE and PBE0 improve the agreement with experiments; however, significant errors remain. On the other hand, RSH and sc-hybrid reproduce the main features of the PE spectrum well, providing the water valence band maximum (VBM) and BEs in agreement with experiments to within about 5%, and the  $\text{NO}_3^-$  IP to within 0.3 eV as compared to experimental data. The overall performance of different DFT calculations is summarized in Fig. 2 using the mean absolute percentage error (MAPE) of the theoretical  $\text{NO}_3^-$  IP, water BEs, and water VBM with respect to the corresponding experimental values. It is shown that the RSH and sc-hybrid functionals exhibit the best performance, yielding errors of 3.9 and 2.3% compared to experiments, respectively. We also note that the relative difference between the ion IP and the water VBM is weakly sensitive to the choice of the density functional, indicating that the error of the ion IP obtained at the conventional PBE level of theory is largely due to the underestimation of the water VBM, as already pointed out in other studies (35, 36).



**Fig. 2. Accuracy of theoretical approaches.** The green boxes represent MAPEs between experiments and DFT calculations, as computed for the NO<sub>3</sub><sup>-</sup> IP, water VBM, and all water BEs. The red and blue boxes correspond to the MAPEs computed only for the NO<sub>3</sub><sup>-</sup> IP and water VBM, as obtained with DFT and G<sub>0</sub>W<sub>0</sub> calculations, respectively. Theoretical calculations include DFT and G<sub>0</sub>W<sub>0</sub> with several exchange-correlation functionals: PBE, HSE, PBE0, RSH, and sc-hybrid functionals (see text).

In addition to DFT calculations, electronic properties of aqueous solutions were investigated using MBPT. Quasi-particle energies, which correspond directly to excitation energies measured in PE spectroscopy, were determined at the G<sub>0</sub>W<sub>0</sub> level of theory using DFT wave functions. The G<sub>0</sub>W<sub>0</sub> calculations were carried out using the scheme recently proposed by Govoni and Galli (22), Nguyen *et al.* (37), and Pham *et al.* (38), which allows for controlled convergence of quasi-particle energies. Only the water VBM and NO<sub>3</sub><sup>-</sup> IP were computed due to the high computational cost of MBPT calculations (see Table 1).

Results presented in Table 1 and Fig. 2 show that G<sub>0</sub>W<sub>0</sub> definitely outperforms DFT when PBE, HSE, and PBE0 wave functions are used. On the other hand, G<sub>0</sub>W<sub>0</sub> calculations based on RSH and sc-hybrid wave functions show only modest improvement compared to DFT. Specifically, the MAPEs with respect to the experiment of NO<sub>3</sub><sup>-</sup> IP and water VBM are 0.7 and 1.3% for G<sub>0</sub>W<sub>0</sub>@RSH and G<sub>0</sub>W<sub>0</sub>@sc-hybrid, respectively, as compared to values of 2.8 and 2.4% obtained with DFT@RSH and DFT@sc-hybrid. Overall, these results indicate that RSH and sc-hybrid functionals provide an accurate and robust description of the electronic properties of liquid water and the NO<sub>3</sub><sup>-</sup> ion within a DFT framework while requiring a much lower computational cost than G<sub>0</sub>W<sub>0</sub> calculations. This conclusion is consistent with a previous study of aqueous solutions with simple ions, that is, Na<sup>+</sup> and Cl<sup>-</sup> (26).

### Theoretical predictions

Having established the level of theory capable of providing a proper description of the electronic properties of the representative NO<sub>3</sub><sup>-</sup> solution, we extended our study to other aqueous solutions to examine the predictive power of first-principles MD coupled with DDH functionals. The moderate computational cost of the hybrid functional calculations allowed us to investigate a relatively large set of ions. In particular, we considered 15 additional solutions containing solvated ions with different atomic species and charge states, which are of interest in studies of electrochemical energy conversion and storage, including NO<sub>2</sub><sup>-</sup>, SO<sub>4</sub><sup>2-</sup>, SO<sub>3</sub><sup>2-</sup>, HCO<sub>3</sub><sup>-</sup>, CO<sub>3</sub><sup>2-</sup>, PO<sub>4</sub><sup>3-</sup>, HPO<sub>4</sub><sup>2-</sup>, ClO<sub>4</sub><sup>-</sup>, ClO<sub>3</sub><sup>-</sup>, ClO<sub>2</sub><sup>-</sup>, and ClO<sup>-</sup> (30, 39–42) and CN<sup>-</sup>, OCN<sup>-</sup>, SCN<sup>-</sup>, and SeCN<sup>-</sup> (43, 44). The structural models of the solutions were generated using

**Table 1. Electronic structure of the NO<sub>3</sub><sup>-</sup> solution.** The water VBM and BEs (1b<sub>1</sub>, 3a<sub>1</sub>, 1b<sub>2</sub>, and 2a<sub>1</sub>) and the NO<sub>3</sub><sup>-</sup> IP were computed using DFT and MBPT at the G<sub>0</sub>W<sub>0</sub> level with different exchange-correlation functionals (see text). All electronic energies (eV) are referred to vacuum using the average potential computed at the same level of theory for water surface models (see Materials and Methods).

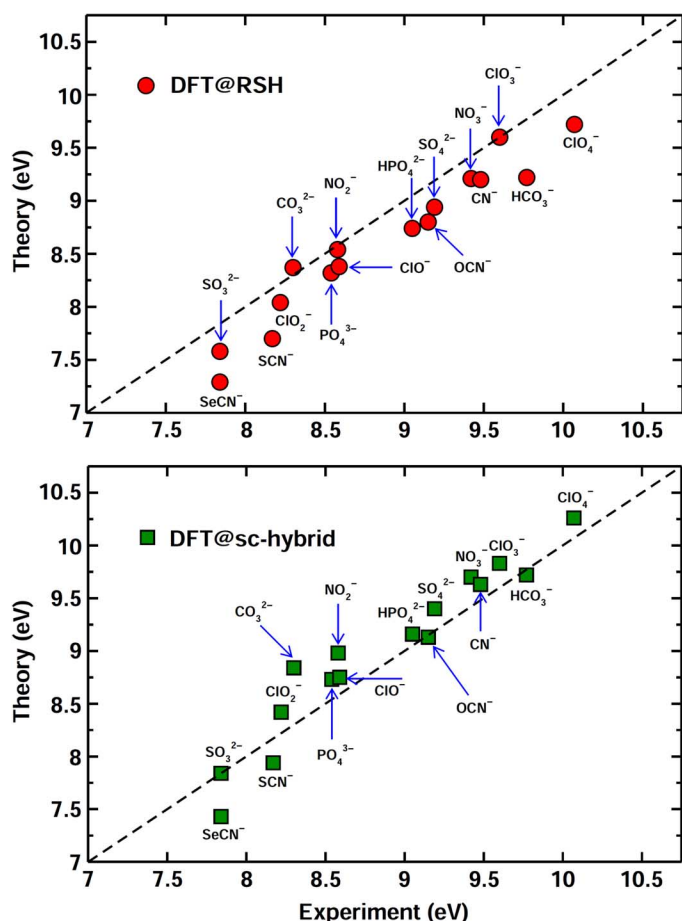
Method	2a <sub>1</sub>	1b <sub>2</sub>	3a <sub>1</sub>	1b <sub>1</sub>	Water VBM	NO <sub>3</sub> <sup>-</sup> IP
DFT@PBE	24.35	12.44	9.11	6.89	5.83	5.37
DFT@HSE	27.20	13.94	10.61	8.27	7.16	6.81
DFT@PBE0	27.59	14.36	11.00	8.69	7.53	7.15
DFT@RSH	30.41	16.49	13.04	10.66	9.57	9.21
DFT@sc-hybrid	32.16	17.01	13.61	11.15	10.08	9.70
G <sub>0</sub> W <sub>0</sub> @PBE	—	—	—	—	8.42	7.72
G <sub>0</sub> W <sub>0</sub> @HSE	—	—	—	—	9.23	8.54
G <sub>0</sub> W <sub>0</sub> @PBE0	—	—	—	—	9.40	8.73
G <sub>0</sub> W <sub>0</sub> @RSH	—	—	—	—	10.02	9.40
G <sub>0</sub> W <sub>0</sub> @sc-hybrid	—	—	—	—	10.10	9.48
Experiment	30.90 <sup>30</sup>	17.41 <sup>31</sup>	13.78 <sup>31</sup>	11.31 <sup>31</sup>	9.90 <sup>30</sup>	9.42

the same simulation protocol used for NO<sub>3</sub><sup>-</sup>. The IPs of solvated ions were computed using DFT with RSH and sc-hybrid functionals, and we compared the theoretical results to PE liquid-jet experiments (see table S1 for details).

Comparison between RSH and sc-hybrid functional calculations and liquid-jet experiments for the IPs of 16 solvated ions is reported in Fig. 3. The overall agreement between theory and experiment is very good, yielding a mean absolute error of 0.14 and 0.11 eV when using RSH and sc-hybrid functionals, respectively. In addition, we found that the discrepancy between theory and experiment is less than 0.5 eV for all ions in a wide energy range of IPs. Compared to RSH, the sc-hybrid functional systematically predicts larger IPs by 0.15 to 0.50 eV, leading to a slightly better overall agreement with experiments. These results provide robust evidence for the predictive power of the DDH functionals in describing the electronic properties of aqueous solutions.

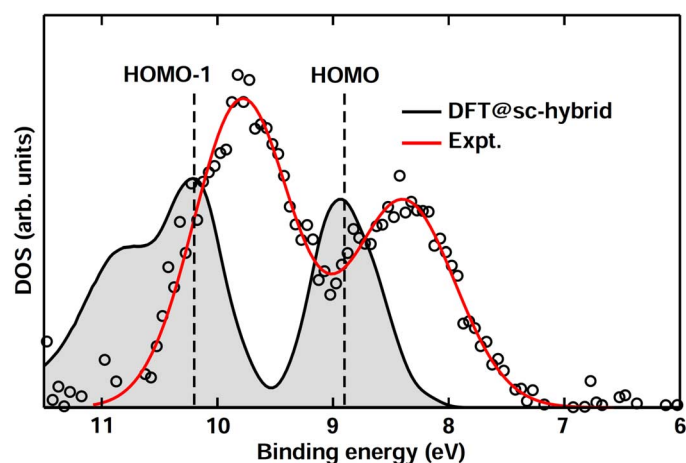
Our study of a wide variety of aqueous solutions also allowed for investigating the ion effects on the electronic properties of liquid water. The latter appears to be slightly perturbed in the presence of ions at the concentration considered here, that is, the water 1b<sub>1</sub> BE and VBM computed for different aqueous solutions show small variations on the order of 0.3 eV. However, we emphasize that ion effects may be more significant for solutions with higher salt concentrations.

To further corroborate the predictive capability of our theoretical framework, we investigated the sensitivity of the results to the details of the simulation protocol. We find that the ion IP is not sensitive to the level of theory used in the MD simulation when comparing semilocal and hybrid functionals at similar temperatures. For example, using simulation trajectories generated with PBE and PBE0, we found that the IP of CO<sub>3</sub><sup>2-</sup> shows a small variation on the order of 0.2 eV (Supplementary Materials). In addition, the use of a fraction of exact exchange  $\alpha$  determined from the electronic dielectric constant  $\epsilon_{\infty}$  of aqueous solutions instead of that of liquid water does not significantly



**Fig. 3. Theoretical predictions.** The IPs of 16 solvated anions computed with DFT and the RSH (circles, top panel) and sc-hybrid (squares, bottom panel) functionals compared with liquid-jet experimental measurements.

affect the reported IPs. This is due to a weak dependence of the refractive index on ion concentration and composition (45), which results in a small change of  $\epsilon_\infty$ . Taking the 1.5 M  $\text{Na}_2\text{CO}_3$  solution as an example, we found that using  $\alpha$  determined from the experimental value of  $\epsilon_\infty = 1.85$  for the solution ( $\epsilon_\infty = 1.78$  for liquid water) resulted in a variation within 0.2 eV of the IP of  $\text{CO}_3^{2-}$ . We also considered finite-size errors in the calculation of ion IPs, which stem from the interaction between periodic images of the ion and the neutralizing background (19). Although accounting for finite-size errors in simulations of aqueous liquids remains a topic of active research (46), finite-size corrections can be estimated reasonably well by following the scheme used for charged defects in semiconductors. Specifically, the leading contribution to the error in defect electronic energy levels is proportional to  $\epsilon^{-1}V^{-1/3}$ , where  $\epsilon$  and  $V$  are the dielectric constant of the semiconductor and the volume of the simulation cell, respectively (19, 47–49). Here, the use of the static dielectric constant of liquid water ( $\epsilon_0 = 78$ ) (50) is appropriate because in our computational framework, we only compute the electronic structure of solvated ions in their ground-state atomic configuration, which is obtained with MD and is thus in thermal equilibrium with the solvent (51). That is, we deduce the ion IPs from the electronic structure of the unperturbed ions in their ground state, instead of computing the IPs as total energy differences between two different oxidation states of the ions. This results in a small finite-size error in the calculated IPs, for example, on the order



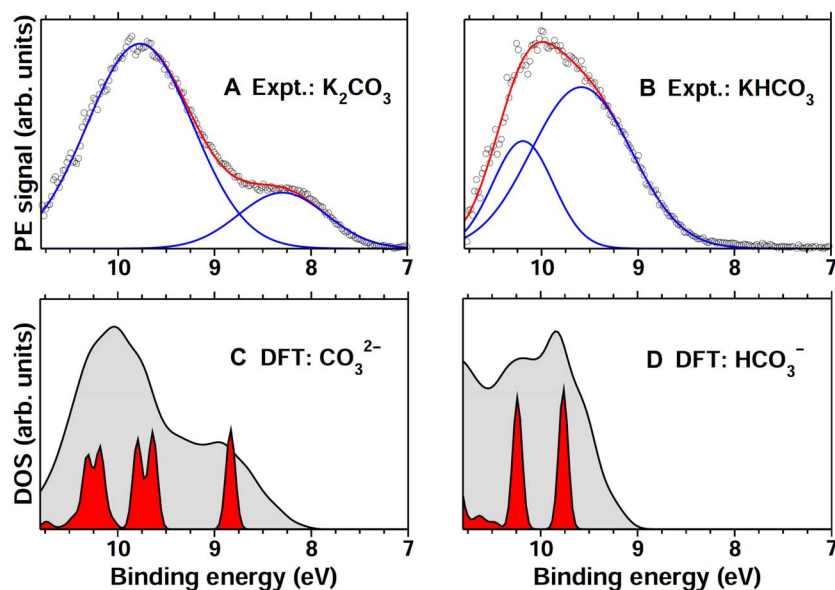
**Fig. 4. PE spectra of the  $\text{PO}_4^{3-}$  solution.** Differential valence PE spectra, measured by liquid-jet techniques, of a 0.4 M  $\text{Na}_3\text{PO}_4$  solution, obtained by subtracting the water spectrum from the raw  $\text{Na}_3\text{PO}_4$  spectrum. The decompose fit of the experimental spectral (open circles) is indicated by the red line. DOS computed with DFT and the sc-hybrid functional for the solvated  $\text{PO}_4^{3-}$  ion is indicated by the gray areas. The anion HOMO and HOMO-1 levels identified from calculations are indicated by the dashed lines. Experimental and theoretical DOS intensities were rescaled with respect to the lower BE peak.

of 0.1 eV for  $\text{NO}_3^-$  for the supercell used here, indicating that finite-size errors in aqueous solutions are largely mitigated by the high static dielectric constant, unlike most solid semiconductors that exhibit much smaller dielectric constants. We note that the second-order correction to the finite-size effects is proportional to  $V^{-1}$ ; this is discussed in more detail by Ayala and Sprik (52) and in the Supplementary Materials.

Finally, we emphasize that at variance with calculations using solvation models (53, 54), DFT with the DDH functionals represents a fully quantum mechanical approach that treats the solutes and solvent on an equal footing. Although the use of solvation models has successfully reproduced experimental IPs for several ions (10, 11, 13), there is growing evidence that a large number of explicit water molecules are needed for a proper description of complex ions (12, 55). To highlight the importance of a quantum mechanical treatment for the liquid environment, we discuss in detail the solvated  $\text{PO}_4^{3-}$  ion that has been theoretically investigated using a hybrid model consisting of a solvated ion cluster and a dielectric continuum (12). In particular, we report in Fig. 4 the  $\text{PO}_4^{3-}$  spectrum obtained from the measurement of a 0.4 M  $\text{Na}_3\text{PO}_4$  solution, which exhibits two peaks at 8.5 and 9.9 eV. Calculations by Pluhařová *et al.* (12) showed that at least two solvation shells of explicit water are required to recover the peak at 8.5 eV that corresponds to the anion HOMO level; however, the higher BE peak in the experimental spectrum was not successfully reproduced. We found that DFT with the sc-hybrid functional not only recovers the double-peak feature in the experimental spectrum but also yields the peak positions in good agreement with experiment (within 0.4 eV) while allowing for identifying the peak at 9.9 eV as the (HOMO-1) level of  $\text{PO}_4^{3-}$ . This finding underscores the need for a fully quantum mechanical approach to obtain a proper description of the electronic properties of complex solvated ions.

### Interpretation of experiments

When coupled with experiments, particularly the liquid-jet technique, the theoretical method presented here offers great promise for a detailed



**Fig. 5. PE spectra of the  $\text{HCO}_3^-$  and  $\text{CO}_3^{2-}$  solutions.** (A and B) Differential valence PE spectra, measured by liquid-jet techniques, of 1 M  $\text{K}_2\text{CO}_3$  (A) and  $\text{KHCO}_3$  (B) aqueous solutions, obtained by subtracting the water spectrum from the raw 1 M  $\text{KHCO}_3$  and  $\text{K}_2\text{CO}_3$  spectra, respectively. The experimental spectra (open circles, Expt.) are decomposed into two Gaussians (blue), with the composite fit shown in red. (C and D) DOS computed with DFT and the sc-hybrid functional for solvated  $\text{CO}_3^{2-}$  and  $\text{HCO}_3^-$  ions are indicated by the gray areas in (C) and (D), respectively. Projected DOS of a representative snapshot [red areas in (C) and (D)] is also presented, showing the positions of different molecular orbitals for each ion [in increasing order of BE, from HOMO to HOMO-4 in (C) and from HOMO to HOMO-1 in (D)].

investigation of complex aqueous solutions, beyond the determination of the ion IPs. In particular, it can be used to interpret the origin of experimental features and elucidate the relationship between chemistry and electronic properties. We present below an example of the  $\text{K}_2\text{CO}_3$  and  $\text{KHCO}_3$  solutions, and we focus our discussion on the results obtained with sc-hybrid functional because it exhibits a slightly better overall performance than the RSH for the ion set presented in the previous section.

The liquid-jet PE spectrum of the  $\text{K}_2\text{CO}_3$  solution (Fig. 5A) exhibits two distinct peaks at 9.77 and 8.3 eV, which may be simply assigned based on the comparison with the  $\text{KHCO}_3$  spectrum (Fig. 5B). In particular, the latter shows a broad peak around 10 eV while lacking one at 8.3 eV. This suggests that the lower BE peak in the  $\text{K}_2\text{CO}_3$  spectrum might be assigned to electrons ejected from  $\text{CO}_3^{2-}$  only, whereas the higher one may arise from the  $\text{HCO}_3^-$  presence in the  $\text{K}_2\text{CO}_3$  solution because of the protonation of  $\text{CO}_3^{2-}$ .

Theoretical calculations provide a more complete and complex interpretation of the PE spectra. Consistent with experiments, the DOS computed for the solvated  $\text{HCO}_3^-$  shows a broad peak around 10 eV (Fig. 5D), thereby supporting the abovementioned interpretation. However, we found that  $\text{CO}_3^{2-}$  also exhibits large DOS in this energy range, which stems from molecular orbitals with higher BEs than the HOMO (Fig. 5C). We therefore conclude that the peak at 8.3 eV in the  $\text{K}_2\text{CO}_3$  spectrum corresponds to the  $\text{CO}_3^{2-}$  HOMO, whereas the one at 9.77 eV has two contributions, that is, from  $\text{CO}_3^{2-}$  molecular orbitals higher in BE than the HOMO and from  $\text{HCO}_3^-$  molecular orbitals. The latter were found to be the HOMO and HOMO-1 levels of  $\text{HCO}_3^-$  based on our theoretical calculations. This example demonstrates that the combined use of first-principles simulations and liquid-jet experiments provides a powerful tool that allows for the identification of chemical species in the solutions and a direct mapping between chemical species and PE signals.

## CONCLUSIONS

In summary, we reported a joint theoretical and experimental study of the electronic properties of aqueous solutions, focused on solvated anions, whose frontier electronic energy levels often play a key role in determining the chemical reactivity of the solutions. We showed that DFT calculations with DDH functionals provide a predictive approach for the study of the electronic properties of aqueous solutions, allowing for a quantitative description of both the solvent PE spectra and the solute electronic energy positions. This level of theory shows similar performance compared to more sophisticated, yet computationally more expensive, MBPT calculations. In addition, we demonstrated that the proposed theoretical technique provides a powerful tool for the interpretation of experimental spectra, allowing for detailed mapping between specific photoemission signals and chemical species. The functionals used here were derived from first-principles methods that treat the solutes and solvents on an equal footing and can be applied to any nonmetallic liquid. Our computational framework paves the way to understanding and engineering the electronic properties of liquid electrolytes for important technologies, such as PEC cells for hydrogen production and ionic liquid for batteries.

## MATERIALS AND METHODS

### MD simulations

All aqueous solutions were simulated using cubic cells consisting of one anion and 63 water molecules with the cell size chosen to yield the experimental water density under ambient conditions, corresponding to a 0.87 M concentration. Counterions were not included, and a positive uniform background charge was used to obtain a neutral system. Born-Oppenheimer molecular dynamics (BOMD) simulations were carried out in the microcanonical ensemble using the Qbox code (<http://qboxcode.org/>) (56), with interatomic forces derived from

DFT using the PBE approximation for the exchange and correlation energy functional (57). The interaction between valence electrons and ionic cores was represented by norm-conserving pseudopotentials (58, 59), and the electronic wave functions were expanded in a plane-wave basis set truncated at a cutoff energy of 85 rydberg (Ry). Hydrogen atoms were replaced with deuterium to increase the allowable time step, which was chosen to be 10 atomic units in all simulations. An elevated temperature of 400 K was used to recover the experimental water structure at  $T = 300$  K (60–62). Analyses of electronic properties were carried out for trajectories collected over 40 ps after a 15-ps equilibration.

The PBE0 simulation of the  $\text{CO}_3^{2-}$  ion was carried out using the same protocol as described above for PBE simulations. The recursive subspace bisection algorithm with a threshold of 0.02 was used (63); this value was chosen based on a previous study of a NaCl solution, showing that the use of the bisection algorithm with a threshold of 0.02 was an excellent approximation to obtain both structural and electronic properties of the solution while allowing for up to a 10-fold acceleration of BOMD simulations with hybrid functionals (64). Electronic structure calculations were performed using a 20-ps NVT (canonical ensemble) trajectory after a 5-ps equilibration run.

### Electronic structure calculations

Calculations of the electronic properties with hybrid functionals were carried out using the Quantum-ESPRESSO (65) and WEST (<http://west-code.org/>) (22) codes. The DOS of each solution was computed using a Gaussian smearing of 0.14 eV and averaged over 100 equally spaced snapshots along the simulation trajectory of 40 ps. The water VBM and ion IPs were computed based on the projected DOS.

MBPT calculations were carried out using the WEST code. The quasi-particle energies  $E_n^{\text{qp}}$  were computed within the  $G_0W_0$  approximation as a first-order correction to the Kohn-Sham (KS) energies  $\epsilon_n$

$$E_n^{\text{qp}} = \epsilon_n + \langle \Psi_n | \Sigma_{G_0W_0}(E_n^{\text{qp}}) | \Psi_n \rangle - \langle \Psi_n | V_{\text{xc}}^{\text{GKS}} | \Psi_n \rangle$$

where  $V_{\text{xc}}^{\text{GKS}}$  is the exchange-correction potential in the generalized KS Hamiltonian and  $\Sigma_{G_0W_0}$  is the self-energy operator computed from the one particle Green's function ( $G_0$ ) and the screened Coulomb interaction ( $W_0$ ). (20).  $G_0W_0$  calculations based on PBE wave functions and eigenvalues were carried out for 20 equally spaced snapshots along the simulation trajectory, whereas those started from hybrid functionals were performed for a single representative snapshot because of their high computational cost compared to semilocal functionals.

Direct comparison between theoretical calculations and liquid-jet experiments requires the evaluation of energy levels relative to vacuum. Absolute DFT and  $G_0W_0$  single-particle energies were determined using the average electrostatic potential computed for water surface models, as discussed in detail in our previous studies (23, 26). This computational scheme is similar to the one used in solid-state physics for the determination of metal work functions and semiconductor band positions (66). It has also been demonstrated to provide electronic band positions of liquid water and simple ions, consistent with those derived from thermodynamic integration techniques (24, 25).

### Liquid-jet experiments

PE measurements were performed at the U41-PGM undulator beamline at the synchrotron radiation facility BESSY II in Berlin, Germany. The liquid microjet with a diameter of 24  $\mu\text{m}$  was injected into the vacuum chamber from a fused silica nozzle. The jet velocity was approx-

imately 40  $\text{ms}^{-1}$ , and the temperature was 6°, except for  $\text{NaH}_2\text{PO}_4$ ,  $\text{Na}_2\text{HPO}_4$ , and  $\text{Na}_3\text{PO}_4$  salt measurements, where we used 20° for solubility reasons. Details of the liquid microjet technique and of the experimental setup have been described previously (67–69). The energy resolution of the U41-PGM beamline was better than 60 meV at a 200-eV photon energy used for the liquid water valence PE measurements. The resolution of the hemispherical energy analyzer, 100 meV at a 10-eV pass energy, was constant with kinetic energy. The small focal size (23  $\mu\text{m} \times 12 \mu\text{m}$ ) of the incident x-ray beam matches the diameter of the liquid jet that leads to a minimal (5%) PE signal contribution from gas-phase water surrounding the jet.

All salt solutions were freshly prepared and immediately injected into the vacuum. Salts were purchased from Sigma-Aldrich (purity, >98%) and not further purified. The counterion in each case was sodium or potassium. All solutions were measured at different concentrations, usually 0.25, 0.5, 1, and 2 M; exceptions are the abovementioned phosphate salts (measured at 0.4 M) and  $\text{Na}_2\text{SO}_4$  (only at 1 M). Valence PE spectra were energy-calibrated against the 1b<sub>1</sub> level of water, which is 11.31 eV (70). Note that a pressure-driven liquid stream in the discharging capillary disrupts an electric double layer created at the interface between the inner wall of the liquid-jet nozzle and the liquid (68, 70, 71). This leads to a streaming potential,  $\Phi_{\text{str}}$ , caused by the charge that is transported by the mobile region of the stream. At a particular electrolyte (salt) concentration (20 to 100 mM),  $\Phi_{\text{str}}$  exactly disappears (70). Studies on halide aqueous solutions, where these  $\Phi_{\text{str}}$  values have been determined, show that previous determinations of solute BEs from several electrolytes and water were typically off by 0.1 to 0.3 eV (70).

### SUPPLEMENTARY MATERIALS

Supplementary material for this article is available at <http://advances.sciencemag.org/cgi/content/full/3/6/e1603210/DC1>

Photoelectron spectra of the  $\text{NaNO}_3$  solution

Calculated  $\text{CO}_3^{2-}$  IPs for different simulation trajectories

IPs of solvated ions

Finite-size effects

fig. S1. Photoelectron spectra of a 1.0 M  $\text{NaNO}_3$  solution measured by the liquid-jet technique.

fig. S2. Radial distribution functions of water molecules computed for  $\text{CO}_3^{2-}$  solutions generated using PBE at 380 K (black) and 400 K (red) and using PBE0 at 380 K (blue).

fig. S3. Radial distribution functions of carbon-water oxygen and carbon-water hydrogen computed for  $\text{CO}_3^{2-}$  solutions generated using PBE at 380 K (black) and 400 K (red) and using PBE0 at 380 K (blue).

fig. S4. DOS computed with DFT and the PBE functional for  $\text{CO}_3^{2-}$  solutions generated with PBE at 400 K (red), PBE at 380 K (blue), and PBE0 at 380 K (green).

table S1. IPs (eV) of selected solvated ions computed using DFT with RSH and sc-hybrid density functionals, compared with PE liquid-jet measurements.

### REFERENCES AND NOTES

1. K. Xu, Nonaqueous liquid electrolytes for lithium-based rechargeable batteries. *Chem. Rev.* **104**, 4303–4418 (2004).
2. P. Simon, Y. Gogotsi, Materials for electrochemical capacitors. *Nat. Mater.* **7**, 845–854 (2008).
3. J. R. McKone, N. S. Lewis, H. B. Gray, Will solar-driven water-splitting devices see the light of day? *Chem. Mater.* **26**, 407–414 (2014).
4. M. G. Walter, E. L. Warren, J. R. McKone, S. W. Boettcher, Q. Mi, E. A. Santori, N. S. Lewis, Solar water splitting cells. *Chem. Rev.* **110**, 6446–6473 (2010).
5. D. V. Esposito, J. B. Baxter, J. John, N. S. Lewis, T. P. Moffat, T. Ogitsu, G. D. O'Neill, T. A. Pham, A. A. Talin, J. M. Velazquez, B. C. Wood, Methods of photoelectrode characterization with high spatial and temporal resolution. *Energy Environ. Sci.* **8**, 2863–2885 (2015).
6. T. A. Pham, Y. Ping, G. Galli, Modelling heterogeneous interfaces for solar water splitting. *Nat. Mater.* **16**, 401–408 (2017).
7. S. P. Ong, O. Andreussi, Y. Wu, N. Marzari, G. Ceder, Electrochemical windows of room-temperature ionic liquids from molecular dynamics and density functional theory calculations. *Chem. Mater.* **23**, 2979–2986 (2011).

8. M. Gauthier, T. J. Carney, A. Grimaud, L. Giordano, N. Pour, H.-H. Chang, D. P. Fenning, S. F. Lux, O. Paschos, C. Bauer, F. Maglia, S. Lupart, P. Lamp, Y. Shao-Horn, Electrode-electrolyte interface in Li-ion batteries: Current understanding and new insights. *J. Phys. Chem. Lett.* **6**, 4653–4672 (2015).
9. J. Cheng, X. Liu, J. VandeVondele, M. Sulprizi, M. Sprik, Redox potentials and acidity constants from density functional theory based molecular dynamics. *Acc. Chem. Res.* **47**, 3522–3529 (2014).
10. B. Jagoda-Cwiklik, P. Slaviček, L. Cwiklik, D. Nolting, B. Winter, P. Jungwirth, Ionization of imidazole in the gas phase, microhydrated environments, and in aqueous solution. *J. Phys. Chem. A* **112**, 3499–3505 (2008).
11. C. A. Schroeder, E. Pluhařová, R. Seidel, W. P. Schroeder, M. Faubel, P. Slaviček, B. Winter, P. Jungwirth, S. E. Bradforth, Oxidation half-reaction of aqueous nucleosides and nucleotides via photoelectron spectroscopy augmented by ab initio calculations. *J. Am. Chem. Soc.* **137**, 201–209 (2015).
12. E. Pluhařová, M. Ončák, R. Seidel, C. Schroeder, W. Schroeder, B. Winter, S. E. Bradforth, P. Jungwirth, P. Slaviček, Transforming anion instability into stability: Contrasting photoionization of three protonation forms of the phosphate ion upon moving into water. *J. Phys. Chem. B* **116**, 13254–13264 (2012).
13. D. Ghosh, A. Roy, R. Seidel, B. Winter, S. Bradforth, A. I. Krylov, First-principle protocol for calculating ionization energies and redox potentials of solvated molecules and ions: Theory and application to aqueous phenol and phenolate. *J. Phys. Chem. B* **116**, 7269–7280 (2012).
14. J. Paier, M. Marsman, K. Hummer, G. Kresse, I. C. Gerber, J. G. Angyan, Screened hybrid density functionals applied to solids. *J. Chem. Phys.* **124**, 154709 (2006).
15. J. H. Skone, M. Govoni, G. Galli, Self-consistent hybrid functional for condensed systems. *Phys. Rev. B* **89**, 195112 (2014).
16. J. H. Skone, M. Govoni, G. Galli, Nonempirical range-separated hybrid functionals for solids and molecules. *Phys. Rev. B* **93**, 235106 (2016).
17. A. Alkauskas, P. Broqvist, F. Devynck, A. Pasquarello, Band offsets at semiconductor-oxide interfaces from hybrid density-functional calculations. *Phys. Rev. Lett.* **101**, 106802 (2008).
18. A. Alkauskas, P. Broqvist, A. Pasquarello, Defect energy levels in density functional calculations: Alignment and band gap problem. *Phys. Rev. Lett.* **101**, 046405 (2008).
19. C. Freysoldt, B. Grabowski, T. Hickel, J. Neugebauer, G. Kresse, A. Janotti, C. G. Van de Walle, First-principles calculations for point defects in solids. *Rev. Mod. Phys.* **86**, 253 (2014).
20. M. S. Hybertsen, S. G. Louie, First-principles theory of quasiparticles: Calculation of band gaps in semiconductors and insulators. *Phys. Rev. Lett.* **55**, 1418–1421 (1985).
21. Y. Ping, D. Rocca, G. Galli, Electronic excitations in light absorbers for photoelectrochemical energy conversion: First principles calculations based on many body perturbation theory. *Chem. Soc. Rev.* **42**, 2437–2469 (2013).
22. M. Govoni, G. Galli, Large scale GW calculations. *J. Chem. Theory Comput.* **11**, 2680–2696 (2015).
23. T. A. Pham, C. Zhang, E. Schwegler, G. Galli, Probing the electronic structure of liquid water with many-body perturbation theory. *Phys. Rev. B* **89**, 060202(R) (2014).
24. D. Opalka, T. A. Pham, M. Sprik, G. Galli, The ionization potential of aqueous hydroxide computed using many-body perturbation theory. *J. Chem. Phys.* **141**, 034501 (2014).
25. D. Opalka, T. A. Pham, M. Sprik, G. Galli, Electronic energy levels and band alignment for aqueous phenol and phenolate from first principles. *J. Phys. Chem. B* **119**, 9651–9660 (2015).
26. A. P. Gaiduk, M. Govoni, R. Seidel, J. H. Skone, B. Winter, G. Galli, Photoelectron spectra of aqueous solutions from first principles. *J. Am. Chem. Soc.* **138**, 6912–6915 (2016).
27. J. Cheng, J. VandeVondele, Calculation of electrochemical energy levels in water using the random phase approximation and a double hybrid functional. *Phys. Rev. Lett.* **116**, 086402 (2016).
28. A. H. England, A. M. Duffin, C. P. Schwartz, J. S. Uejio, D. Prendergast, R. J. Saykally, On the hydration and hydrolysis of carbon dioxide. *Chem. Phys. Lett.* **514**, 187–195 (2011).
29. J. W. Smith, R. K. Lam, O. Shih, A. M. Rizzuto, D. Prendergast, R. J. Saykally, Properties of aqueous nitrate and nitrite from x-ray absorption spectroscopy. *J. Chem. Phys.* **143**, 084503 (2015).
30. C. Zhong, Y. Deng, W. Hu, J. Qiao, L. Zhang, J. Zhang, A review of electrolyte materials and compositions for electrochemical supercapacitors. *Chem. Soc. Rev.* **44**, 7484–7539 (2015).
31. C. R. Cox, J. Z. Lee, D. G. Nocera, T. Buonassisi, Ten-percent solar-to-fuel conversion with nonprecious materials. *Proc. Natl. Acad. Sci. U.S.A.* **111**, 14057–14061 (2014).
32. B. Winter, R. Weber, W. Widdra, M. Dittmar, M. Faubel, I. V. Hertel, Full valence band photoemission from liquid water using EUV synchrotron radiation. *J. Phys. Chem. A* **108**, 2625–2632 (2004).
33. B. Winter, R. Weber, I. V. Hertel, M. Faubel, P. Jungwirth, E. C. Brown, S. E. Bradforth, Electron binding energies of aqueous alkali and halide ions: EUV photoelectron spectroscopy of liquid solutions and combined ab initio and molecular dynamics calculations. *J. Am. Chem. Soc.* **127**, 7203–7214 (2005).
34. J. Heyd, J. E. Peralta, G. E. Scuseria, R. L. Martin, Energy band gaps and lattice parameters evaluated with the Heyd-Scuseria-Ernzerhof screened hybrid functional. *J. Chem. Phys.* **123**, 174101 (2005).
35. C. Adriaanse, J. Cheng, V. Chau, M. Sulprizi, J. VandeVondele, M. Sprik, Aqueous redox chemistry and the electronic band structure of liquid water. *J. Phys. Chem. Lett.* **3**, 3411–3415 (2012).
36. F. Ambrosio, G. Miceli, A. Pasquarello, Redox levels in aqueous solution: Effect of van der Waals interactions and hybrid functionals. *J. Chem. Phys.* **143**, 244508 (2016).
37. H.-V. Nguyen, T. A. Pham, D. Rocca, G. Galli, Improving accuracy and efficiency of calculations of photoemission spectra within the many-body perturbation theory. *Phys. Rev. B* **85**, 081101(R) (2012).
38. T. A. Pham, H.-V. Nguyen, D. Rocca, G. Galli, GW calculations using the spectral decomposition of the dielectric matrix: Verification, validation, and comparison of methods. *Phys. Rev. B* **87**, 155148 (2013).
39. M. W. Kanan, D. G. Nocera, In situ formation of an oxygen-evolving catalyst in neutral water containing phosphate and  $\text{Co}^{2+}$ . *Science* **321**, 1072–1075 (2008).
40. R. Saito, Y. Miseki, K. Sayama, Highly efficient photoelectrochemical water splitting using a thin film photoanode of  $\text{BiVO}_4/\text{SnO}_2/\text{WO}_3$  multi-composite in a carbonate electrolyte. *Chem. Commun.* **48**, 3833–3835 (2012).
41. Y. Zhang, C. Sun, P. Lu, K. Li, S. Song, D. Xue, Crystallization design of  $\text{MnO}_2$  towards better supercapacitance. *CrystEngComm* **14**, 5892–5897 (2012).
42. Z. Li, W. Luo, M. Zhang, J. Feng, Z. Zou, Photoelectrochemical cells for solar hydrogen production: Current state of promising photoelectrodes, methods to improve their properties, and outlook. *Energy Environ. Sci.* **6**, 347–370 (2013).
43. P. Wang, S. M. Zakeeruddin, J.-E. Moser, R. Humphry-Baker, M. Grätzel, A solvent-free,  $\text{SeCN}^-/\text{(SeCN)}_3^-$  based ionic liquid electrolyte for high-efficiency dye-sensitized nanocrystalline solar cells. *J. Am. Chem. Soc.* **126**, 7164–7165 (2004).
44. A. Hagfeldt, G. Boschloo, L. Sun, L. Kloo, H. Pettersson, Dye-sensitized solar cells. *Chem. Rev.* **110**, 6595–6663 (2010).
45. A. V. Wolf, *Aqueous Solutions and Body Fluids* (Harper and Row, 1966).
46. X. Liu, J. Cheng, M. Sprik, Aqueous transition-metal cations as impurities in a wide gap oxide: The  $\text{Cu}^{2+}/\text{Cu}^+$  and  $\text{Ag}^{2+}/\text{Ag}^+$  redox couples revisited. *J. Phys. Chem. B* **119**, 1152–1163 (2015).
47. S. Lany, A. Zunger, Assessment of correction methods for the band-gap problem and for finite-size effects in supercell defect calculations: Case studies for ZnO and GaAs. *Phys. Rev. B* **78**, 235104 (2008).
48. H.-P. Komsa, T. T. Rantala, A. Pasquarello, Finite-size supercell correction schemes for charged defect calculations. *Phys. Rev. B* **86**, 045112 (2012).
49. W. Chen, A. Pasquarello, Correspondence of defect energy levels in hybrid density functional theory and many-body perturbation theory. *Phys. Rev. B* **88**, 115104 (2013).
50. M. Sharma, R. Resta, R. Car, Dipolar correlations and the dielectric permittivity of water. *Phys. Rev. Lett.* **98**, 247401 (2007).
51. P. Delahay, Photoelectron emission spectroscopy of aqueous solutions. *Acc. Chem. Res.* **15**, 40–45 (1982).
52. R. Ayala, M. Sprik, A classical point charge model study of system size dependence of oxidation and reorganization free energies in aqueous solution. *J. Phys. Chem. B* **112**, 257–269 (2008).
53. J. Tomasi, B. Mennucci, R. Cammi, Quantum mechanical continuum solvation models. *Chem. Rev.* **105**, 2999–3094 (2005).
54. C. J. Cramer, D. G. Truhlar, A universal approach to solvation modeling. *Acc. Chem. Res.* **41**, 760–768 (2008).
55. D. Yepes, R. Seidel, B. Winter, J. Blumberger, P. Jaque, Photoemission spectra and density functional theory calculations of 3d transition metal-aqua complexes (Ti–Cu) in aqueous solution. *J. Phys. Chem. B* **118**, 6850–6863 (2014).
56. F. Gygi, Architecture of Qbox: A scalable first-principles molecular dynamics code. *IBM J. Res. Dev.* **52**, 137–144 (2008).
57. J. P. Perdew, K. Burke, M. Ernzerhof, Generalized gradient approximation made simple. *Phys. Rev. Lett.* **77**, 3865–3868 (1996).
58. D. Vanderbilt, Optimally smooth norm-conserving pseudopotentials. *Phys. Rev. B* **32**, 8412–8415 (1985).
59. M. Schlupf, F. Gygi, Optimization algorithm for the generation of ONCV pseudopotentials. *Comput. Phys. Commun.* **196**, 36–44 (2015).
60. J. C. Grossman, E. Schwegler, E. W. Draeger, F. Gygi, G. Galli, Towards an assessment of the accuracy of density functional theory for first principles simulations of water. *J. Chem. Phys.* **120**, 300 (2004).
61. E. Schwegler, J. C. Grossman, F. Gygi, G. Galli, Towards an assessment of the accuracy of density functional theory for first principles simulations of water. II. *J. Chem. Phys.* **121**, 5400–5409 (2004).
62. T. A. Pham, T. Ogitsu, E. Y. Lau, E. Schwegler, Structure and dynamics of aqueous solutions from PBE-based first-principles molecular dynamics simulations. *J. Chem. Phys.* **145**, 154501 (2016).
63. F. Gygi, Compact representations of Kohn-Sham invariant subspaces. *Phys. Rev. Lett.* **102**, 166406 (2009).
64. A. P. Gaiduk, C. Zhang, F. Gygi, G. Galli, Structural and electronic properties of aqueous NaCl solutions from ab initio molecular dynamics simulations with hybrid density functionals. *Chem. Phys. Lett.* **604**, 89–96 (2014).

65. P. Giannozzi, S. Baroni, N. Bonini, M. Calandra, R. Car, C. Cavazzoni, D. Ceresoli, G. L. Chiarotti, M. Cococcioni, I. Dabo, A. Dal Corso, S. de Gironcoli, S. Fabris, G. Fratesi, R. Gebauer, U. Gerstmann, C. Gougousis, A. Kokalj, M. Lazzeri, L. Martin-Samos, N. Marzari, F. Mauri, R. Mazzarello, S. Paolini, A. Pasquarello, L. Paulatto, C. Sbraccia, S. Scandolo, G. Sclauzero, A. P. Seitsonen, A. Smogunov, P. Umari, R. M. Wentzcovitch, QUANTUM ESPRESSO: A modular and open-source software project for quantum simulations of materials. *J. Phys. Condens. Matter* **21**, 395502 (2009).
66. T. A. Pham, D. Lee, E. Schwegler, G. Galli, Interfacial effects on the band edges of functionalized Si surfaces in liquid water. *J. Am. Chem. Soc.* **136**, 17071–17077 (2014).
67. B. Winter, Liquid microjet for photoelectron spectroscopy. *Nucl. Instr. Methods A* **601**, 139–150 (2009).
68. B. Winter, M. Faubel, Photoemission from liquid aqueous solutions. *Chem. Rev.* **106**, 1176–1211 (2006).
69. R. Seidel, S. Thurmer, B. Winter, Photoelectron spectroscopy meets aqueous solution: Studies from a vacuum liquid microjet. *J. Phys. Chem. Lett.* **2**, 633–641 (2011).
70. N. Kurahashi, S. Karashima, Y. Tang, T. Horio, B. Abulimiti, Y.-I. Suzuki, Y. Ogi, M. Oura, T. Suzuki, Photoelectron spectroscopy of aqueous solutions: Streaming potentials of NaX (X= Cl, Br, and I) solutions and electron binding energies of liquid water and X<sup>-</sup>. *J. Chem. Phys.* **140**, 174506 (2014).
71. M. Faubel, B. Steiner, J. P. Toennies, Photoelectron spectroscopy of liquid water, some alcohols, and pure nonane in free micro jets. *J. Chem. Phys.* **106**, 9013 (1997).

**Acknowledgments:** We thank A. Gaiduk, J. Skone, D. Prendergast, C. Van de Walle, J. Varley, V. Lordi, and B. Winter for fruitful discussions. We also thank the BESSY II staff for their assistance during the beam times. **Funding:** Part of this work was performed under the auspices of the U.S. Department of Energy (DOE) at Lawrence Livermore National Laboratory

under contract DE-AC52-07A27344. T.A.P. acknowledges support from the Lawrence Fellowship. The software used in this project was supported by Midwest Integrated Center for Computational Materials (MICCoM) (M.G. and G.G.), as part of the Computational Materials Sciences Program funded by the DOE, Office of Science, Basic Energy Sciences, Materials Sciences and Engineering Division (5J-30161-0010A). R.S. gratefully acknowledges support from the Deutsche Forschungsgemeinschaft (SE 2253/3-1). S.E.B. acknowledges support from the NSF (CHE-1301465). E.S. acknowledges support from the DOE/Basic Energy Sciences (grant no. DE-SC0008938). Computational resources were from the Lawrence Livermore National Laboratory Institutional Computing Grand Challenge Program and the Argonne Leadership Computing Facility (contract DE-AC02-06CH11357). Author contributions: T.A.P. and G.G. designed the research. T.A.P. performed most of the calculations. M.G. implemented the GW MBPT and the DDH functionals in the WEST code. R.S. performed liquid-jet experiments. All authors contributed to the analysis and discussion of the data and the writing of the manuscript. **Competing interests:** The authors declare that they have no competing interests. **Data and materials availability:** All data needed to evaluate the conclusions in the paper are present in the paper and/or the Supplementary Materials. Additional data related to this paper may be requested from the authors.

Submitted 16 December 2016

Accepted 28 April 2017

Published 23 June 2017

10.1126/sciadv.1603210

**Citation:** T. A. Pham, M. Govoni, R. Seidel, S. E. Bradforth, E. Schwegler, G. Galli, Electronic structure of aqueous solutions: Bridging the gap between theory and experiments. *Sci. Adv.* **3**, e1603210 (2017).



## Electronic structure of aqueous solutions: Bridging the gap between theory and experiments

Tuan Anh Pham, Marco Govoni, Robert Seidel, Stephen E. Bradforth, Eric Schwegler, and Giulia Galli

*Sci. Adv.*, **3** (6), e1603210.

DOI: 10.1126/sciadv.1603210

### View the article online

<https://www.science.org/doi/10.1126/sciadv.1603210>

### Permissions

<https://www.science.org/help/reprints-and-permissions>

Use of this article is subject to the [Terms of service](#)

---

*Science Advances* (ISSN 2375-2548) is published by the American Association for the Advancement of Science, 1200 New York Avenue NW, Washington, DC 20005. The title *Science Advances* is a registered trademark of AAAS. Copyright © 2017 The Authors, some rights reserved; exclusive licensee American Association for the Advancement of Science. No claim to original U.S. Government Works. Distributed under a Creative Commons Attribution NonCommercial License 4.0 (CC BY-NC).