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Evidence of catalase mimetic activity in ce(3+)/ce(4+) doped bioactive glasses / Nicolini, Valentina; Gambuzzi, Elisa; Malavasi, Gianluca; Menabue, Ledi; Menziani, Maria Cristina; Lusvardi, Gigliola; Pedone, Alfonso; Benedetti, Francesco; Luches, Paola; D'Addato, Sergio; Valeri, Sergio. - In: JOURNAL OF PHYSICAL CHEMISTRY. B, CONDENSED MATTER, MATERIALS, SURFACES, INTERFACES & BIOPHYSICAL. - ISSN 1520-6106. - ELETTRONICO. - 119:10(2015), pp. 4009-4019. [10.1021/jp511737b]

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25/04/2024 23:50

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Article

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J. Phys. Chem. B, **Just Accepted Manuscript** • DOI: 10.1021/jp511737b • Publication Date (Web): 24 Feb 2015Downloaded from <http://pubs.acs.org> on February 26, 2015**Just Accepted**

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Evidence of Catalase Mimetic Activity in Ce³⁺/Ce⁴⁺ Doped Bioactive Glasses

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Abstract.

The ability of Ce-containing bioactive glasses to inhibit oxidative stress in terms of reduction of hydrogen peroxide, by mimicking the catalase enzyme activity is demonstrated here for the first time.

The antioxidant properties of three bioactive glasses containing an increasing amount of CeO₂ have been evaluated by following the degradation of hydrogen peroxide with time after immersion in H₂O₂ aqueous solutions with different concentration. XPS and UV-Vis measurements allowed us to determine the Ce³⁺/Ce⁴⁺ ratio in the bulk and on the glass surface, and to correlate it with the ability of the samples to show catalase mimetic activity. Interestingly, we have found that the bioactive glass with composition 23.2Na₂O-25.7CaO-43.4SiO₂-2.4P₂O₅ 5.3CeO₂ immersed in 0.1 M H₂O₂ aqueous solution is able to degrade 90% of it in one week. The reduction in bioactivity of the glasses with increasing CeO₂ content is here rationalized in terms of a lower amount of phosphate groups available for the hydroxyapatite layer formation, after binding with cerium ions. In fact, classical molecular dynamics simulations revealed that the addition of CeO₂ leads to the formation of cerium phosphate rich regions. The formation of an insoluble CePO₄ crystalline phase is also observed by XRD analysis after thermal treatment of the glass samples.

Keywords. Cerium, biomaterials, antioxidant properties, molecular dynamics, XRD, UV.

1. Introduction

In the field of biomaterial implantation, the so-called surgical stress response is a well-defined physiological mechanism that involves, during and after surgical procedures, the activation of inflammatory, endocrine, metabolic and immunologic mediators.¹⁻³

Surgical stress also includes the occurrence of oxidative stress, with production of reactive oxygen or nitrogen species that may overwhelm the defense systems of the organism. It has been demonstrated that the administration of antioxidants results in improved organ function, shortened convalescence, reduced morbidity and mortality occurring in the surgical stress response.^{4,5}

In this context, biomaterials that can contrast the effects of oxidative stress and inhibit excessive reactive oxygen species generation in a sustained manner may be a useful tool for therapies that target these medical problems.

Among biomaterials, bioactive glasses are widely used in bone defect reparation because they spontaneously bond and integrate with both soft and hard (bone) tissues in the living body.⁶ In fact, when bioactive glasses are implanted and in contact with biological fluids, a rapid formation of a thin hydroxycarbonate-apatite layer occurs on the surface of the material, enhancing living bone tissue regeneration.⁷ Since the early 1970s, when the first Bioglass[®] (45S5 glass: 24.5% Na₂O-24.5% CaO-45.0% SiO₂-6% P₂O₅ by weight)⁸ was synthesized, glasses of many different compositions have been studied to improve the material properties.⁹⁻¹³ For example, by varying the chemical nature and concentration of metal oxides, new important biological properties have been added, tailoring the glass to specific clinical applications.¹⁴⁻²¹

1 The design of bioactive glasses able to prevent oxidative stress after implantation would shorten the
2 convalescence and reduce the amount of anti-inflammatory medications administered to patients. More-
3 over, the use of bioactive glasses able to act as antioxidant can be helpful in the management of osteopo-
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10 In the recent years, it has been shown that nanoparticles of cerium oxide,²³⁻²⁸ RuO₂,²⁹ Fe₃O₄,²⁹ and
11 Co₃O₄³⁰ are able to catalyze the decomposition of hydrogen peroxide (H₂O₂) into water and oxygen
12 mimicking the catalase enzyme action. The role of the latter is very important, since it protects cells
13 from oxidative stress caused by reactive oxygen species.^{31,32}

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20 The ability of these nanoparticles to act against oxidative stress is closely related to the ability of the
21 metal to present two oxidation states. For example, nanoparticles of ceria, CeO₂ (nanoceria), present
22 oxygen vacancies associated to the presence of Ce³⁺ on the surface. Their surface exhibits a high
23 Ce³⁺/Ce⁴⁺ ratio, whereas bulk CeO₂ is characterized by a low Ce³⁺/Ce⁴⁺ ratio.^{30,33} Thus, the redox prop-
24 erties of nanoceria determine its ability to protect tissues against oxidative stress.³⁴⁻³⁷

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31 In this scenario, we will investigate the possibility to imprint catalase mimetic activity to Ce-doped
32 bioactive glasses.³⁸ Characterization of magnetic, optical and structural properties will be carried out
33 both with experimental and computational methods. The latter approach will allow us to get a detailed
34 description of the medium-range structural arrangement (~5 Å) of glasses and to correlate it with mac-
35 roscopic properties. To the best of our knowledge, the ability of Ce-containing bioactive glasses to in-
36 hibit oxidative stress in terms of reduction of hydrogen peroxide will be demonstrated here for the first
37 time. A relationship between antioxidant and bioactivity properties will also be discussed.

38 39 40 41 42 43 44 45 46 47 48 **2. Experimental section**

49 50 51 **2.1 Preparation of glass powders.**

1 Three different glasses containing 1.2%, 3.6% and 5.3% of CeO₂ (molar composition) were synthe-
2 sized *via* melting procedure starting from the composition of the 45S5 Bioglass[®] (named hereafter BG),
3 following the procedure describe in ref.³⁸
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7 The nominal and effective molar composition of each sample is reported in Table 1. A BG sample was
8 also prepared as reference. Samples were prepared by mixing reagent grade SiO₂, Na₂CO₃, CaCO₃,
9 Na₃PO₄·12H₂O and CeO₂ in a glass beaker. Then each batch was put into a platinum crucible and melt-
10 ed in an electric oven. Heating ramp was set to 15°C/min up to 1000°C and 8°C/min up to 1350°C.
11 Samples were maintained at this temperature for 2 hours, to ensure an optimal melting and mixing of all
12 the oxides, and finally quenched at room temperature on a graphite plate. All the synthesized specimens
13 were milled and sieved to obtain glass powders with particles in the range of 250-500µm.
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23 2.2 Characterization of glasses

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27 *ICP measurements.* The effective composition of each sample as quenched (AQ) was determined after
28 dissolution by the Perkin Elmer Optima 4200DV inductive coupled plasma – optical emission spec-
29 trometer (ICP-OES), equipped with ultrasonic nebulizer CETAC useful for trace elements.
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34 *UV-Vis spectroscopy.* To gain a qualitative estimation of the Ce⁴⁺ contained in the samples, UV-Vis
35 analyses were performed on the as quenched powders. Spectra were acquired by using the HP8452 UV-
36 Vis spectrometer in the 200-850 nm range for all samples. The diffuse reflectance technique with a Ba-
37 SO₄ plate as reflectance standard was used.
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44 *X-Ray Diffraction (XRD) characterization.* Powders of the doped glasses were crystallized at 660°C
45 for 2 hours, then they were analyzed in the (2Θ) 10°-50° range, by means of X-ray diffraction apparatus
46 (the X'Pert PRO-PANAnalytical) equipped with Ni-filtered Cu K_α radiation (λ = 1.54060 Å).
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51 *X-Ray Photoelectron Spectroscopy (XPS).* XPS was used to obtain quantitative information on the rel-
52 ative amount of Ce³⁺ and Ce⁴⁺ on the surface of the samples. XPS spectra were collected at normal
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1 emission using a hemispherical electron analyzer and Al K_{α} photons as the exciting probe. Due to the
2 insulating nature of the samples, the XPS spectra were affected by charging effects, which resulted in a
3 shift of the binding energy of the photoemission peaks. The shift ranges between 4.6 and 5.5 eV in the
4 different samples. To account for this effect, the binding energy of each Ce 3d XPS spectrum was cor-
5 rected by subtracting the difference between the binding energy of the 1s peak of adventitious carbon
6 measured on the sample and the reference value of 285 eV.³⁹ In spite of the low concentration of Ce
7 contained on the surface of the investigated samples, we managed to measure Ce 3d XPS spectra with
8 an acceptable signal to noise ratio using long acquisition times (approximately 1 hour for the Ce 3d
9 spectrum). The spectra were fit following a well established procedure, using three spin-orbit split dou-
10 blets related to Ce⁴⁺ ionic species and two doublets related to Ce³⁺ ionic species.⁴⁰⁻⁴³ The spin-orbit
11 splitting and branching ratio of each component and the differences in binding energy between the com-
12 ponents were fixed to the values in ref. ⁴⁰⁻⁴³ The Lorentzian and Gaussian width of the individual com-
13 ponents were kept as close as possible to the values used in previous works on cerium oxide films with
14 different degree of oxidation.⁴¹⁻⁴³ In the fitting of the different spectra shown in this work the area of the
15 peaks and a quadratic background were the only free parameters, while the binding energy, spin orbit
16 splitting, branching ratio, widths and step of the Shirley type background for all components were fixed.
17 The concentration of each ionic species was obtained from the relative weight of the area of its compo-
18 nents compared to the total area of all Ce 3d components. It has to be noted that the accuracy of the con-
19 centration of the two Ce ionic species obtained in this way is not very high, but the considerations made
20 in the following are mainly based on the evolution of their values in the different samples/conditions,
21 which are considered to be reliable. ⁴⁰⁻⁴³ Given the limited probing depth of the XPS technique, the in-
22 formation obtained can be related only to the first few nanometers below the surface of the investigated
23 samples.

1 *Molecular dynamic simulation.* Classical MD simulations of the BG, BG_1.2Ce, BG_3.6Ce and
2 BG_5.3Ce glasses were performed by means of the DL_POLY[®] package⁴⁴ employing a well-established
3 melt-quench computational protocol.⁴⁵
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7 In this approach, an initial random configuration containing about 10,000 atoms enclosed in a period-
8 ic cubic box with the experimental density is melted at 3200 K and then cooled down to 300 K at a
9 nominal cooling rate of 5 K ps⁻¹. The resulting glass structures have been subjected to a final NVT
10 trajectory of 0.3 ns; and the structural analysis was performed on 1001 configurations sampled at
11 regular intervals during the last 0.2 ns of MD trajectory. The Ce⁴⁺/Ce³⁺ ratio of each glass com-
12 position was fixed accordingly to experimental findings. The interatomic forces acting between the
13 constituting ions have been described by using a shell model force field able to accurately reproduce the
14 short and medium range structure of multicomponent oxide glasses.⁴⁶⁻⁵³ In this model, cations are repre-
15 sented by rigid points bearing full formal charges (Si⁴⁺, P⁵⁺, Ca²⁺, Ce³⁺, Ce⁴⁺ etc...), whereas anions are
16 split in a massless shell of charge -Y and a massive core of charge Z (Z+Y is the total atomic charge; -2
17 for O ions), which are coupled by a harmonic spring potential. The oxygen shell interacts with the Si⁴⁺,
18 P⁵⁺, Ce³⁺, Ce⁴⁺, Ca²⁺ and Na⁺ cations through a Buckingham term, and coulomb forces act between all
19 species bearing a formal charge, in the Born Mayer model. Three-body screened harmonic potentials
20 have been employed to control the O-Si-O and O-P-O angles, for favoring tetrahedral coordination.
21 Therefore, in the shell model the polarizability of oxygen is straightforwardly included, and it is also
22 environment dependent due to the forces acting on the oxygen shell by the other ions. This allows for a
23 better glass structure relaxation during quenching and leads to glasses with improved medium-
24 range structures with respect to those generated by using the rigid ion model.⁵⁴ A detailed de-
25 scription of the force-field functional forms as well as the complete list of parameters is reported
26 in refs.^{46,47,50}
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2.3 Catalase mimetic activity tests

In order to verify whether glasses containing CeO₂ show catalase mimetic activity, the powder samples were soaked, under continuous stirring, in solutions of H₂O₂ at concentrations 1 M and 0.1 M.³⁴ The soaking time was set to 1, 2, 4 hours, 1, 4 and 7 days. In all the samples, a constant glass mass/solution volume ratio of 5 mg/ml was maintained. After soaking, the specimens were filtered; the powders were dried overnight at 60°C, whereas the solutions were titrated by KMnO₄ to determine the residual H₂O₂. Before titration, pH of each solution was measured and the release of cerium from the specimens was determined by inductive coupled plasma – mass spectrometer (ICP-MS).

The powders were analyzed by UV-Vis spectroscopy (range 250-650 nm) and XPS in order to obtain qualitative and quantitative information about the Ce³⁺/Ce⁴⁺ ratio.

3. Results

3.1 Glass characterization.

ICP measurements. The effective compositions obtained by ICP analysis are reported in Table 1, together with the nominal ones. Notwithstanding small discrepancies are observed between the nominal and effective compositions, the ratios between the base constituent of the Ce-doped BGs are very similar to the ratios in the original Bioglass[®] 45S5.

UV-Vis spectroscopy. To gather qualitative information about the oxidation state of cerium contained in the glasses, UV-Vis measurements were performed. CeO₂ and Ce(NO₃)₃·6H₂O were used as references to precisely identify the absorption wavelengths of Ce³⁺ and Ce⁴⁺ respectively. Figure 1 shows that Ce(NO₃)₃·6H₂O present a broad peak between 250-350 nm (Figure 1), whereas CeO₂ shows two peaks, one in the range of 250-300 nm and one around 380 nm, in good agreement with previous stud-

ies.⁵⁵⁻⁵⁷ By considering that CeO₂ always carries a small portion of Ce₂O₃ on the surface, it is possible to assign the peak at 250-300 nm to Ce³⁺ and the peak at 350-400 nm at Ce⁴⁺.

The BG sample shows a peak at 270 nm, that is the characteristic absorption wavelength of soda-lime-silicate glass.⁵⁸ Spectra of Ce-doped glasses show very broad peaks between 250-400 nm; thus, both Ce³⁺ and Ce⁴⁺ should be present in our specimens. Furthermore, the intensity of absorption between 400 and 600 nm is directly proportional to the content of CeO₂ in the sample composition, suggesting an increase in the quantity of Ce⁴⁺.

XRD characterization. Previous investigations^{59,60} revealed that after heat treatment the BG sample crystallizes through the formation of a cyclo-silicate (Na₂CaSi₂O₆, JCPDS 77-2189) and a mixed Na-Ca ortho-phosphate (β-NaCaPO₄, JCPDS 76-1456). Figure 2 reports the XRD diffraction patterns before and after heat treatment of the Ce-containing BG samples investigated here. The patterns show peaks characteristic of three different phases: Na₂CaSi₂O₆ (JCPDS 77-2189), CePO₄ (JCPDS 32-199) and CeO₂ (JCPDS 34-0394). The presence of CePO₄ (JCPDS 32-199) and CeO₂ (JCPDS 34-0394) phases is strictly related to the Ce-content in the glasses, in fact the intensity of the peaks attributed to these phases increases passing from 1.2% CeO₂ to 5.3% CeO₂, however in BG_1.2Ce crystallized sample only the peak attributed to CeO₂ was observed. Interestingly, Ce-containing glasses do not present the crystalline phase β-NaCaPO₄ detected in our previous work in BG sample in addition to Na₂CaSi₂O₆ crystals⁶¹, suggesting that the PO₄³⁻ structural units tend to attract cerium ions instead of sodium and calcium and favors the formation of CePO₄ instead of β-NaCaPO₄.

XPS. Figure 3 shows the Ce 3d XPS spectra measured on BG_5.3Ce, BG_3.6Ce and BG_1.2Ce AQ sample after Shirley-type background subtraction. All of the spectra in Figure 3 contain mainly features attributed to Ce³⁺, although the individual peak around 916.5 eV, attributed to Ce⁴⁺, has a non-negligible intensity in all samples. This evidence confirms that cerium contained in the glasses is always present in

1 both its oxidation states, as already observed by UV-Vis and magnetic susceptibility measurements. The
2 fitting of the Ce 3d XPS spectra allowed a quantitative evaluation of the ratio between the two Ce ionic
3 species on the glass surface. The $\text{Ce}^{3+}/\text{Ce}^{4+}$ ratio obtained for the AQ glass surface is 3.2 ± 0.2 for
4 BG_5.3Ce, 2.8 ± 0.3 for BG_3.6Ce and 3.3 ± 0.5 for BG_1.2Ce. As expected, the $\text{Ce}^{3+}/\text{Ce}^{4+}$ ratio does
5 not change significantly in samples with different Ce concentration.
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11 *Molecular dynamic simulations.* An important structural feature related to the dissolution of ions in
12 the physiological fluids is their coordination numbers $\text{CN}_{\text{X-Y}}$, that is, the number of neighbors of species
13 Y in the first coordination shell of a given atom X. Table 2 reports the $\text{CN}_{\text{X-O}}$ of all the cations in the
14 investigated glasses. As expected, both silicon and phosphorus ions are four-coordinated and their coordination numbers are not influenced by the addition of Ce ions.
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24 On average, Ce^{3+} is coordinated by 6.5-6.6 non-bridging oxygens (NBOs) in all glasses, whereas Ce^{4+}
25 is surrounded by about 7.0 NBOs, in agreement with previous MD investigations^{50,62,63} and EXAFS
26 measurements^{64,65} on phosphate glasses. The coordination of cerium by NBOs only means that Ce acts
27 as modifier cation that strongly competes for NBOs coordination with Na and Ca ions.
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34 The latter are both coordinated by about 5.8 oxygens in the BG glass in agreement with previous investigations,⁶⁶⁻⁷⁰ but whereas the CN of Ca ions remains constant, the one of Na ions slightly increases
35 to 6 with the CeO_2 content. The number of NBOs and bridging oxygen (BOs) atoms coordinated to Ca
36 remains essentially constant in the four compositions, whereas a depletion of BO and an increment of
37 NBO are observed around Na cations upon CeO_2 addition.
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46 To investigate the effect of CeO_2 on the glass polymerization, the Q^n distributions (Q^n stands for 4-
47 coordinated network former ion bonded to n bridging oxygens) of Si and P extracted from the MD trajectory for the BG, BG_1.2Ce, BG_3.6Ce and BG_5.3Ce glasses are listed in Table 3.
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1 The cation-cation coordination numbers (CN_{X-Y}) computed by integrating the first peak of the corre-
2 sponding pair distribution functions provides insights on the medium range order and chemical disorder
3 of these glasses. Table 2 shows that the addition of CeO_2 oxide causes an increment in the number of
4 Na ions in the second coordination sphere of Si: from 4.48, for BG glass, to 5.85 for BG_5.3Ce. At the
5 same time, the number of Na surrounding P decreases from 6.05, for BG, to about 5.4-5.7 for Ce-
6 containing compositions. The number of Ca ions around Si and P cations is less dependent on the addi-
7 tion of CeO_2 , though the second coordination sphere of P in BG_5.3Ce presents a lower number of Ca
8 ions with respect to other glasses.
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10 These data suggest that the Ce atoms tend to satisfy their own coordination requirements by using the
11 NBOs belonging to PO_4^{3-} groups. Therefore, when inserted into the BG glass, Ce atoms push Na^+ ions
12 away from P environment toward the silica network, which in turn depolymerizes in order to provide
13 NBOs to Na^+ cations following the usual reaction $Si-O-Si + Na_2O \rightarrow 2SiONa$. The preference of Ce
14 ions for phosphate groups is also witnessed by the comparison between the P/Si ratio around Ce ions in
15 the MD models and the nominal ratio given by the glass composition (Table 2): the P/Si ratio in the sec-
16 ond coordination sphere of Ce^{3+} ions is about 10 times higher than the nominal one (0.11). Figure 4
17 shows that P environment in BG glass (Figure 4.a) is rich in Na and Ca, whereas Ce ions replace part of
18 Na and Ca in BG_5.3Ce (Figure 4.b). The structural differences are evident also in terms of silica net-
19 work depolymerization: BG glass (Figure 4.c) presents a more polymerized structure, where both Ca
20 and Na are allocated close to phosphate tetrahedra, while BG_5.3Ce glass (Figure 4.d) exhibits a highly
21 depolymerized network with a prevalence of Na ions in Si closest surroundings.
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23 The $Q^n(Si)$ distributions for the four compositions and the corresponding network connectivity, NC,
24 (average number of BOs per network-forming atom) are quite different to each other. Although, the Si
25 distribution is always dominated by Q^2 chain species, a significant amount of chain-terminators, Q^1 , and
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1 branched, Q^3 , species are also present in all glasses. The addition of CeO_2 leads to a more fragmented
2 network with the NC monotonically decreasing from 2.08 for the BG glass to 1.74 for BG_5.3Ce glass.
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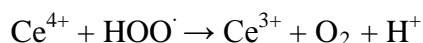
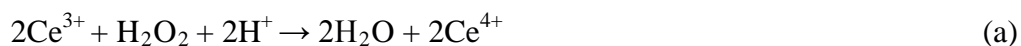
5 As for phosphorus, more than 76% of its ions are found in an orthophosphate (Q^0) environment in the
6 BG glass, but the amount of Q^0 species rapidly increases to 89% after the addition of only 5.3 mol% of
7 CeO_2 . The opposite trend is observed for $P(Q^1)$ sites, while the scant populations of $P(Q^2)$ sites does not
8 allow to speculate on its variation. However, it is evident that Ce has a strong impact on the medium
9 range order of these bioactive glasses.
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16 **3.2 Catalase mimetic activity: in vitro tests and characterization after treatment.**

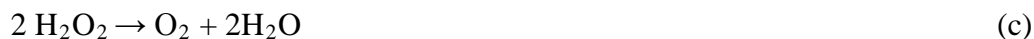
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20 *Titration with $KMnO_4$.* Figure 5 reports the residual H_2O_2 concentrations after different soaking times
21 of the three Ce-containing samples in the original solutions with concentration of 0.1 M and 1.0 M. Both
22 trends highlight that the degradation of H_2O_2 is enhanced for glasses containing more CeO_2 . The results
23 show that the H_2O_2 degradations become apparent after 24 hours and proceed with the reaction time. In
24 particular, the most noticeable results were achieved with BG_3.6Ce and BG_5.3Ce samples soaked in
25 H_2O_2 solution for 168 hours with a reduction of around 64% and 90% of H_2O_2 concentration respec-
26 tively starting from H_2O_2 0.1 M and a reduction of around 17% and 35% respectively starting from
27 H_2O_2 1 M.
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39 Stoichiometric calculations (not shown here) were carried out to determine the ratio between moles of
40 H_2O_2 decomposed after 168 hours (7 days) and moles of cerium contained in the samples. The resulting
41 H_2O_2/Ce ratios for BG_3.6Ce are 58 for H_2O_2 1 M and 22 for H_2O_2 0.1 M, while for BG_5.3Ce they
42 are 80 and 23, respectively. These ratios are underestimated since they were calculated by considering
43 all the amount of cerium contained in the samples, neglecting that only the Ce ions on the surface
44 should be able to interact with the hydrogen peroxide. This demonstrates that the amount of H_2O_2 de-
45 stroyed is in large excess with respect to the quantity of cerium in agreement with a catalytic behavior of
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1 the metal; indeed Ce^{3+} is able to reduce H_2O_2 to H_2O and, at the same time, Ce^{4+} is able to oxidize it
2 first to HOO^\cdot and then to O_2 . Starting from these considerations, we confirm the model of reaction
3 mechanism for the dismutation of hydrogen peroxide by cerium oxide nanoparticles proposed by Celar-
4 do at al.⁷¹ and summarized in the following:
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Summing reactions a) and b) the result is reaction c):



that is the same reaction through which the catalase enzyme destroys the hydrogen peroxide.⁷²

38 *UV-Vis spectroscopy.* The UV-Vis spectra performed over the soaked samples at different time have
39 been compared with the dried powders spectra in Figure 6. The sample without Ce showed an absorb-
40 ance at λ lower than 350 nm. This is characteristic of soda-lime silicate glasses⁷³ while the Ce-
41 containing glasses started to adsorb at higher λ (<500 nm). It is possible to observe that the peaks of
42 soaked samples shift to higher wavelength (λ) with respect to the AQ samples, that is, toward the λ of
43 Ce^{4+} adsorption (380-400 nm). Thus, it is reasonable to assume that during the soaking in hydrogen per-
44 oxide, the Ce^{3+} ions in the samples convert partially to the Ce^{4+} state. Moreover, the intensity of the ab-
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1 sorption at 168 hours is often lower than the intensity at 96 hours, possibly indicating that some Ce^{4+}
2 ions revert to Ce^{3+} due to the catalase mimetic activity of the glasses.
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5 *ICP-MS measurements.* ICP-MS measurements were performed over each filtered solution to gather
6 information over the cerium released in solution during the soaking. As expected, samples with greater
7 content in Ce release more metal and the concentration of Ce in solution falls in the range between 20
8 ppb and 1.5 ppm for all the samples. Furthermore, the greatest concentrations were obtained when
9 glasses were soaked in H_2O_2 1 M, which is more aggressive towards the specimens. Also the greatest
10 value of cerium in solution (1.5 ppm) obtained is not enough to justify the high abatement of H_2O_2
11 reached by soaking the samples, hence cerium ions on the glass surfaces should play a main role in the
12 reaction with the hydrogen peroxide in solution. This hypothesis is confirmed by the fact that no de-
13 composition of H_2O_2 is detected after 7 days for a solution of 1.5 ppm of Ce-salts (with different
14 $\text{Ce}^{3+}/\text{Ce}^{4+}$ ratio) in H_2O_2 1 M. It is important to highlight that since the release of cerium ions in the en-
15 vironment is ten times lower than the LD_{50} reported in literature for cerium ions (15 ppm) the bioactive
16 glasses investigated should not lead to toxicity problems.⁷⁴ For the sake of completeness, we reported
17 also the value of LD_{50} of cerium oxide, which is 1000 ppm, according to OSHA and ANSI.⁷⁵ Further-
18 more, Zhang et al. reported in a recent study⁷⁶ that concentration up to 40 ppm of nanoceria protect the
19 bone marrow stromal cells (BMSC) from the oxidative stress, increasing their viability when H_2O_2 so-
20 lution is added to the environment.
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43 *pH-measurement.* The pH was also monitored during the in vitro tests. An increment from 6.1 to 9.2
44 was observed after 7 days for all the samples in 0.1 M H_2O_2 solution, whereas the pH value increased to
45 10.7 after 7 days for all the samples in 1 M H_2O_2 solution. This increment is a well-known phenomenon
46 due to the exchange between Na^+ at the glass surface with H^+ from the solution occurring in the first
47 step of the Hench's bioactivity mechanism.^{7,77}
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XPS. The analysis of the Ce 3d spectra allowed obtaining quantitative information on the evolution of $\text{Ce}^{3+}/\text{Ce}^{4+}$ ratio in the BG_5.3Ce glass after different soaking times. Figure 7a and b report the Ce 3d spectra for the 1 M and 0.1 M H_2O_2 solutions after selected soaking times and their fitting. A clear evolution of the Ce 3d shape with soaking time can be observed. The full set of values of the $\text{Ce}^{3+}/\text{Ce}^{4+}$ ratio obtained from the fitting at the different soaking times is reported in Figure 7c. A decrease of the intensity of Ce 3d signal was observed for soaking times longer than 1 hour in 1 M H_2O_2 solution, resulting in a lower signal-to-noise ratio in the spectra acquired after 1 day soaking in 1 M H_2O_2 in spite of the relatively long acquisition time. After 7 days for both H_2O_2 concentrations, the Ce 3d signal was almost suppressed. The Ce surface concentration decrease after long soaking times is ascribed to the formation of precipitate substances as Ca-phosphate phases on the glass surface. The $\text{Ce}^{3+}/\text{Ce}^{4+}$ ratio on the glass surface decreases for short soaking times in the 1 M H_2O_2 solution and it subsequently increases, reaching a value close to 1 for longer times. In the 0.1 M H_2O_2 solution the $\text{Ce}^{3+}/\text{Ce}^{4+}$ decrease is more significant and it persists for longer times than in the 1 M solution.

4. Discussion

The characterization carried out on Ce-doped BGs confirms that cerium is present in the two common oxidation states $\text{Ce}^{3+}/\text{Ce}^{4+}$, both in the bulk and on the surface. XPS analysis determined a $\text{Ce}^{3+}/\text{Ce}^{4+}$ ratio in the range 1.6 - 3.1 for the studied glasses.

In a recent work, Pirmohamed *et al.*²³ compared the catalase mimetic activity of two kinds of Ce NPs with different $\text{Ce}^{3+}/\text{Ce}^{4+}$ ratios (6-26) and found the best results for systems with low ratios.

Our results are in agreement with these findings. In fact, we found that all glasses are able to reduce the H_2O_2 concentration. In particular, the BG_Ce5.3 glass presents the best catalytic rate, since it is able to degrade 31.7 μmol of H_2O_2 per min. For example, a sample of 250 mg was able to decompose in 168 h a number of H_2O_2 moles equals to 80 times the moles of cerium contained in the whole bulk sample.

1 This catalytic rate is higher than that found by Pirmohamed *et al.* (2.71 nmol of H₂O₂ per min)²³ proba-
2 bly because the tests were performed using different concentrations of both H₂O₂ (10 mM as maximum
3 value in the previous work) and CeO₂ nanoparticles.
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7 The ability of the glasses to act as catalysts is also highlighted by the trend of Ce³⁺/Ce⁴⁺ after the soak-
8 ing in H₂O₂ solution reported in Figure 7a. The ratio shows a great variation at short time and it be-
9 comes more or less constant at longer time, suggesting a similar velocity of the two proposed reactions
10 a) and b) at longer times. Moreover, at short soaking times (1-4 hours) a significant color change from
11 light yellow to deep yellow-orange was observed indicating that Ce³⁺ (colorless) on the glass surface
12 was oxidized to Ce⁴⁺ (yellow-orange)⁷⁸ by H₂O₂. The color of the sample remained the same in the rest
13 of the observations. Therefore, it can be hypothesized that at short times reaction a) occurs at the glass
14 surface, and, since the specific surface area of the sample is low, this process causes a low decomposi-
15 tion of H₂O₂. Instead, at longer times (>1 day) the catalytic cycle of reaction a) + b) starts, and the in-
16 crement of specific surface area,⁷⁹ due to the partial dissolution of the glass, increases the rate of H₂O₂
17 decomposition since more Ce³⁺ and Ce⁴⁺ ions become available.
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34 Interestingly, the glasses are active in a wide range of pH, from 6 to 9.2 and 10.7 as a function of
35 H₂O₂ concentration and time. This is very important because it is the range of pH commonly reached in
36 the fluids surrounding the implanted biomaterial.⁸⁰
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41 The XRD analysis and MD results suggest the absence of CeO₂ nanocrystals in the glass matrix. Both
42 techniques revealed that in the glass structure Ce ions and phosphate groups tend to aggregate forming
43 Ce-phosphate rich zones. In view of this results we can ascribe the catalase mimetic activity to the sim-
44 ultaneous presence on the material surface of Ce³⁺ and Ce⁴⁺ sites, opening a new door on the develop-
45 ment of inorganic materials exhibiting redox state-dependent mimetic enzyme activity.
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1 From one hand, the increasing of CeO₂ content in the glass compositions enhances anti-oxidant ability,
2 but from another hand, the bioactivity decreases (delay of hydroxyapatite formation) as reported by Le-
3 onelli *et al.*³⁸ The computational simulation of glass structures allowed us to relate this phenomenon
4 with the medium-range order of the matrix. In fact, Ce atoms tend to allocate close to phosphate do-
5 mains, preventing their release from the glass network, thus making them less available for the hydrox-
6 yapatite layer formation, or involving them in the formation of a stable, insoluble CePO₄ crystalline
7 phase. Moreover, the introduction of CeO₂ in these glasses avoids the formation of calcium phosphate
8 domains that can act as crystallization nuclei able to speed up the crystallization of Ca₃(PO₄)₂ on the
9 glass surface and the formation of hydroxyapatite and could influence ions release.^{81,82,83} Therefore a
10 good compromise between a high bioactivity level and an efficient catalase mimetic activity requires a
11 proper adjustment of cerium oxide content in the glasses. An alternative way to obtain both good bioac-
12 tivity and efficient catalytic activity could be the synthesis of the glasses through the sol-gel method.⁸⁴
13 Indeed, in a previous work Shruti *et al.* have verified that bioactive glasses obtained *via* sol-gel and
14 doped with 3.5% mol cerium oxide do not show a substantial difference in bioactivity, compared to the
15 undoped glass.²⁰ This is likely due to the fact that porous sol-gel glasses show higher bioactivity with
16 respect to the glasses obtained by melting. Thus, it is possible that adding an amount greater than 3.5%
17 to the sol-gel glasses, they could show both bioactivity and antioxidant properties.
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48 5. Conclusions

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In conclusion, we have demonstrated for the first time that it is possible to confer catalase mimetic activity to bioactive glasses by adding CeO₂. This property is related to the simultaneous presence of Ce³⁺ and Ce⁴⁺ both in the bulk structure and on the surface with a ratio close to 3.

The structural information gained by XRD measurements and MD simulations showed that the introduction of cerium has a strong impact on the medium range order of the glasses investigated, favoring the formation of Ce-phosphate insoluble phases that increase the chemical durability. Thus, the employment of CeO₂-doped bioactive glasses could become a valid alternative to the addition of ceria nanoparticles to biomaterials.

Acknowledgment

This work was supported by a grant from the University of Modena and Reggio Emilia entitled ‘The role of cerium oxidation state in bioactive glasses used as biomaterials of 3rd generation’. Support by the COST Action CM1104 ‘Reducible oxide chemistry, structure and functions’ is also acknowledged.

Supporting Information Available

Magnetic susceptibility data collected over the AQ samples. Table S1: The magnetic moment (μ_{eff}) of as quenched (AQ) samples. This material is available free of charge *via* the Internet at <http://pubs.acs.org/>.

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Tables**Table 1.** Nominal and effective molar compositions (%) of the synthesized glasses obtained by three different measurements (\pm std. dev.).

Nominal	BG	BG_1.2Ce	BG_3.6Ce	BG_5.3Ce
CeO ₂	-	1.2	3.6	5.3
SiO ₂	46.2	45.6	42.5	43.4
Na ₂ O	24.3	24.0	24.4	23.2
CaO	26.9	26.6	27.0	25.7
P ₂ O ₅	2.6	2.6	2.5	2.4
Effective				
CeO ₂	-	1.0 \pm 0.1	3.2 \pm 0.2	5.1 \pm 0.4
SiO ₂	47.8 \pm 1.5	44.5 \pm 1.6	41.4 \pm 2.0	40.8 \pm 2.5
Na ₂ O	23.5 \pm 2.2	27.4 \pm 2.4	29.0 \pm 2.8	30.1 \pm 2.9
CaO	26.0 \pm 2.4	24.6 \pm 2.3	24.1 \pm 2.2	21.9 \pm 2.2
P ₂ O ₅	2.6 \pm 0.2	2.5 \pm 0.2	2.3 \pm 0.2	2.1 \pm 0.2

Table 2. Average coordination numbers for the first and second coordination shells; ratio of the number of P and Si around Ce^{3+} $((\text{P}/\text{Si})_{\text{Ce}^{3+}})$, Ce^{4+} $((\text{P}/\text{Si})_{\text{Ce}^{4+}})$ and in the glass structure $((\text{P}/\text{Si})_{\text{GLASS}})$.

$\text{CN}_{\text{X-O}}$	BG	BG_1.2Ce	BG_3.6Ce	BG_5.3Ce
First coordination shell ($\text{CN}_{\text{X-O}}$)				
Si-O	4.00	4.00	4.00	4.00
P-O	4.00	4.00	4.00	4.00
Ca-O	5.76	5.82	5.94	5.86
Ca-BO	0.47	0.46	0.54	0.45
Ca-NBO	5.29	5.31	5.40	5.41
Na-O	5.81	5.36	6.02	6.01
Na-BO	1.35	1.36	1.25	1.20
Na-NBO	4.46	4.55	4.77	4.81
Ce^{3+} -NBO*	---	6.53	6.62	6.57
Ce^{4+} -NBO*	---	7.00	7.08	6.98
Second coordination shell ($\text{CN}_{\text{X-Y}}$)				
Si-Na	4.48	5.01	5.00	5.85
Si-Ca	2.93	2.95	2.98	2.89
P-Na	6.05	5.76	5.39	5.68
P-Ca	3.63	3.68	3.68	3.30
P/Si				
$(\text{P}/\text{Si})_{\text{Ce}^{3+}}$	---	1.20	1.07	1.12
$(\text{P}/\text{Si})_{\text{Ce}^{4+}}$	---	0.25	1.37	1.09
$(\text{P}/\text{Si})_{\text{GLASS}}$	----	0.11	0.11	0.11

(*) Ce atoms are coordinated by NBOs only.

Table 3. Q^n distributions and network connectivity, NC (average number of BOs coordinating network forming atoms), of P and Si network former cations, calculated for the studied glasses by MD simulations.

	BG	BG_1.2Ce	BG_3.6Ce	BG_5.3Ce
$Q^n(\text{Si})$ speciation (pop%)				
n=0	0.9	1.2	2.2	3.2
n=1	19.2	22.5	29.2	35.0
n=2	51.9	52.9	51.8	46.9
n=3	25.9	21.9	15.9	13.9
n=4	2.0	1.5	1.0	1.0
NC	2.09	2.00	1.84	1.74
$Q^n(\text{P})$ speciation (pop%)				
n=0	76.7	77.7	85.3	89.4
n=1	22.2	21.2	14.7	10.1
n=2	1.1	1.1	0.0	0.6
NC	0.24	0.23	0.15	0.11

Figures

Figure 1. UV-Vis spectra of glass samples and standard substances containing Ce^{3+} (cerium nitrate hexa-hydrate) and Ce^{4+} (cerium dioxide).

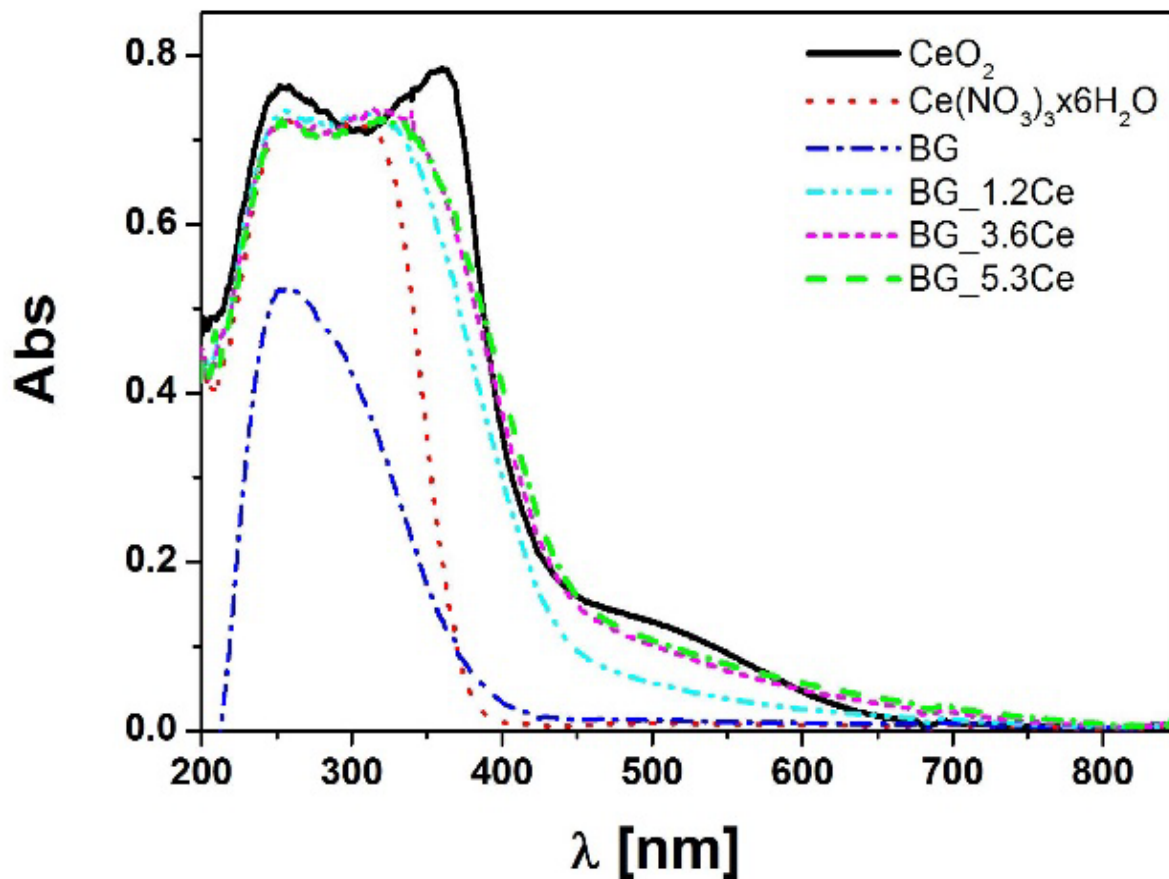


Figure 2. XRD spectra of BG_1.2Ce and BG_3.6Ce crystallized samples at 660°C for 2 hours. XRD spectra of BG_5.3Ce before (AQ) and after crystallization at 660°C for 2 hours. (the XRD spectrum of BG crystallized is reported in reference⁶⁰).

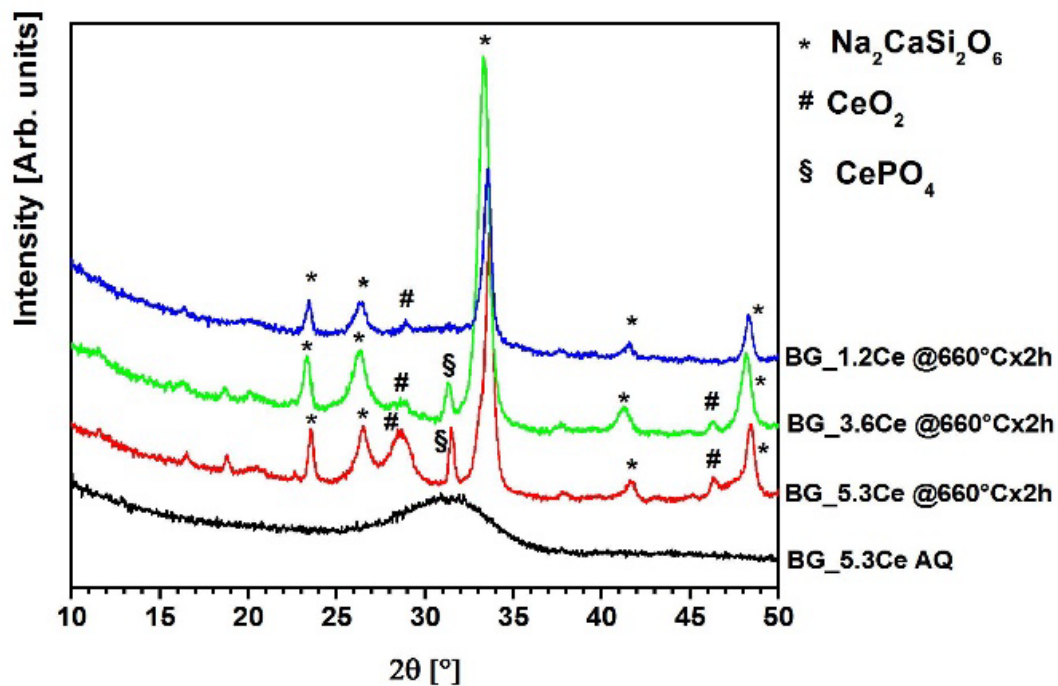


Figure 3. Ce 3d XPS spectra for BG_5.3Ce, BG_3.6Ce and BG_1.2Ce samples (red lines). The spectra are shown after Shirley-type background subtraction. For each peak the fitting curve is also shown (black lines).

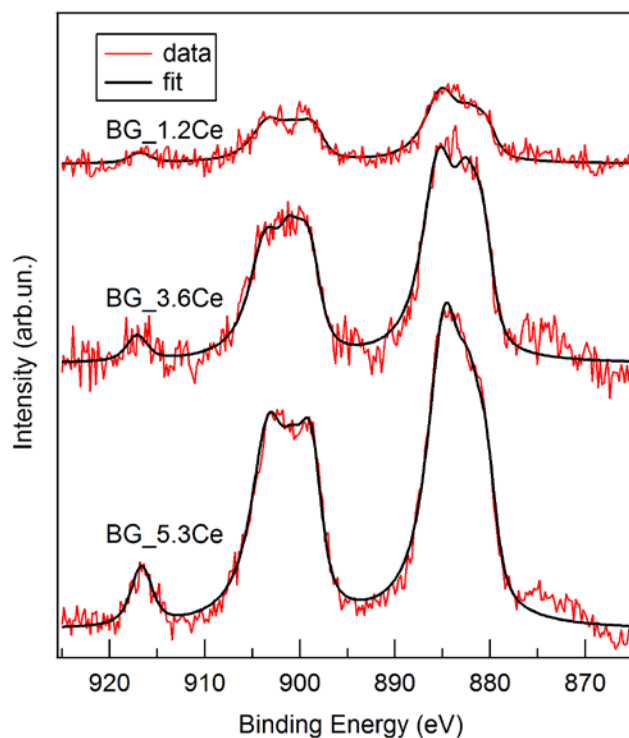


Figure 4. Phosphorous environment in a) BG glass and b) BG_5.3Ce glass; Silicate network polymerization and its environment in c) BG glass and d) BG_5.3Ce glass. Blue and dark yellow tetrahedra are P and Si, respectively. Red, light blue, yellow and green spheres are O, Na, Ca and Ce, respectively.

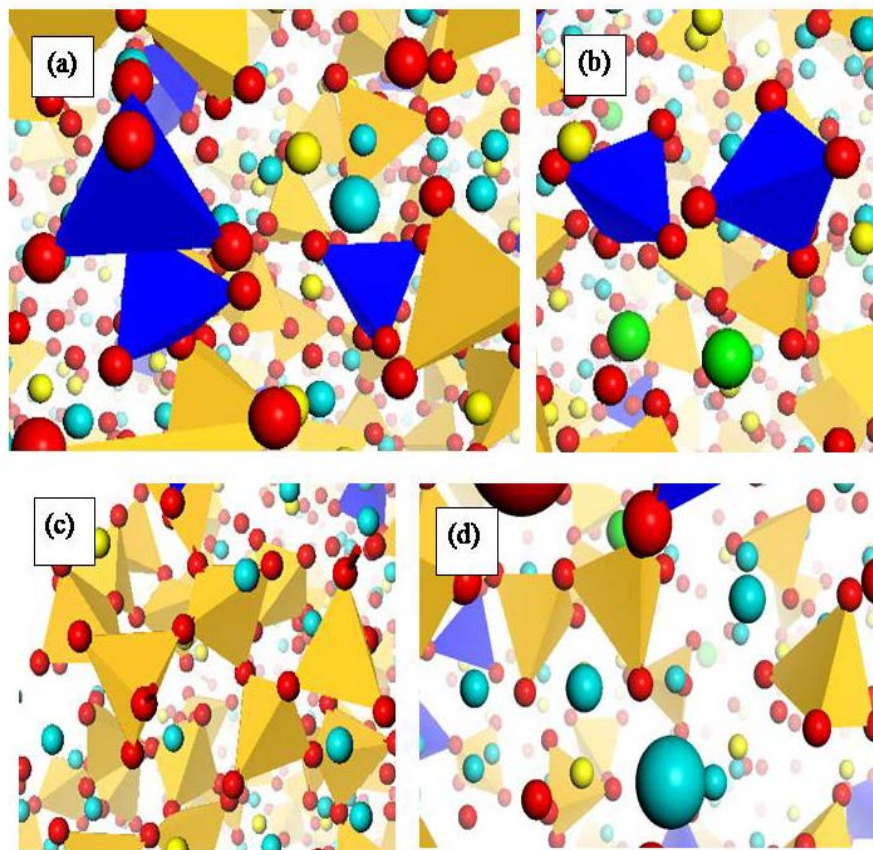


Figure 5. Degradation of (a) H_2O_2 0.1 M and (b) H_2O_2 1 M after soaking of the glasses doped with CeO_2 , BG is reported as control. The lines serve only to guide the eyes.

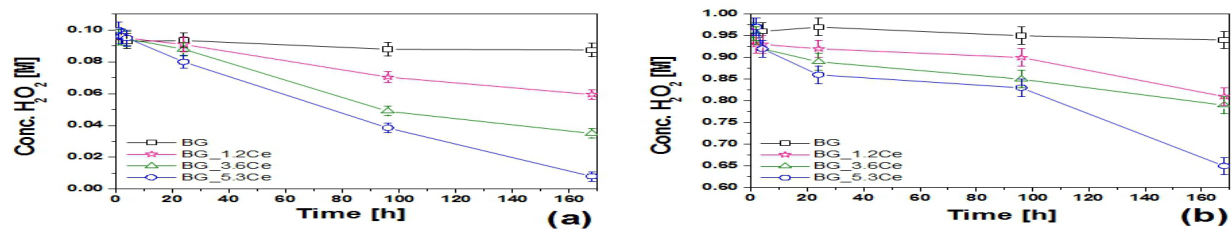


Figure 6. UV-Vis spectra of samples before (AQ) and after soaking in 0.1 M (left) and 1 M (right) H_2O_2 solution at different times.

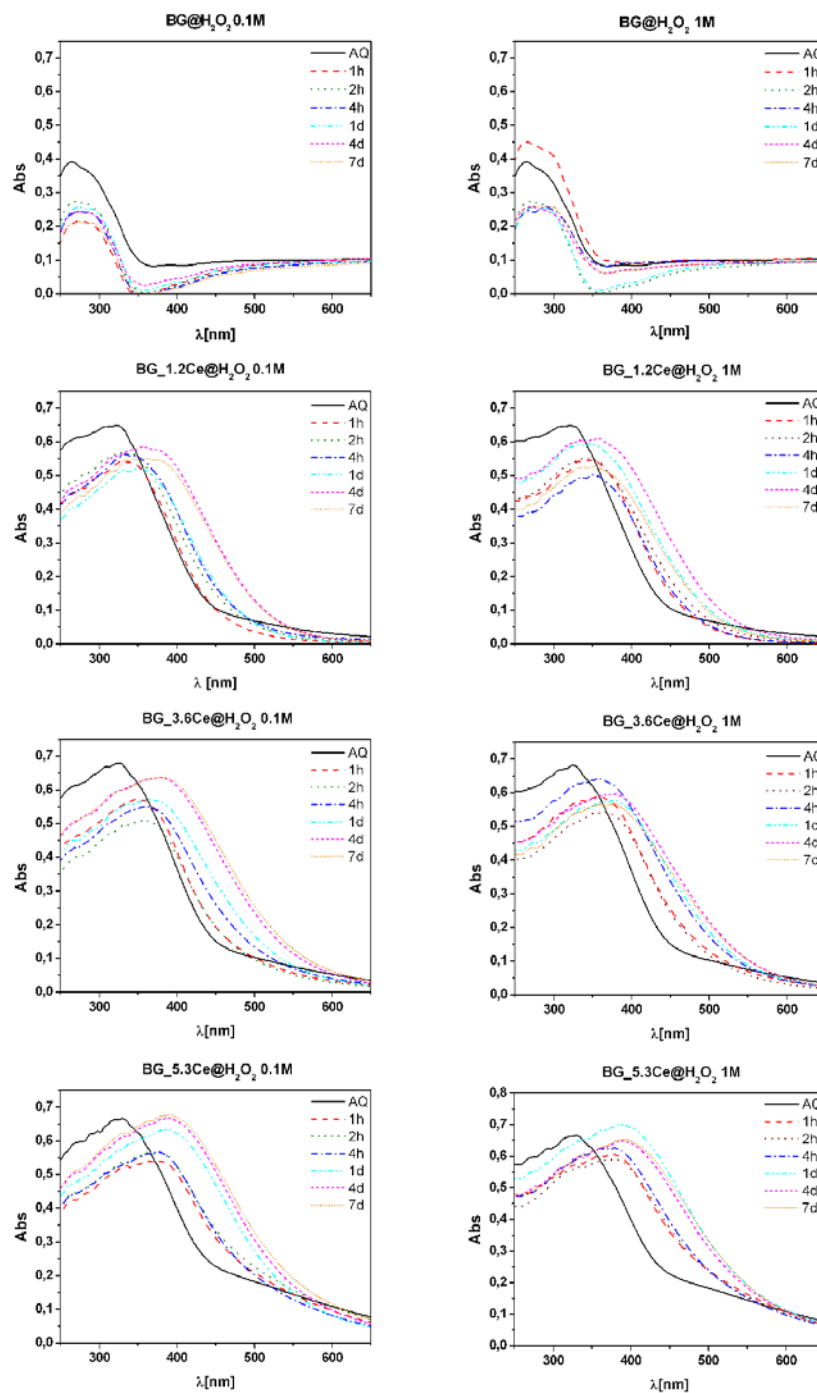


Figure 7. Ce 3d XPS spectra for BG_5.3Ce (red lines) and fitting curves (black lines) after different selected soaking times in 0.1 M (a) and 1 M (b) H_2O_2 solution. The spectra have been normalized in intensity to compare their shape. (c) $\text{Ce}^{3+}/\text{Ce}^{4+}$ ratios obtained from the fitting of Ce 3d XPS spectra as a function of soaking time in the 0.1 M (red dots) and 1 M (blue triangles) H_2O_2 solution; the lines serve only to guide the eyes.

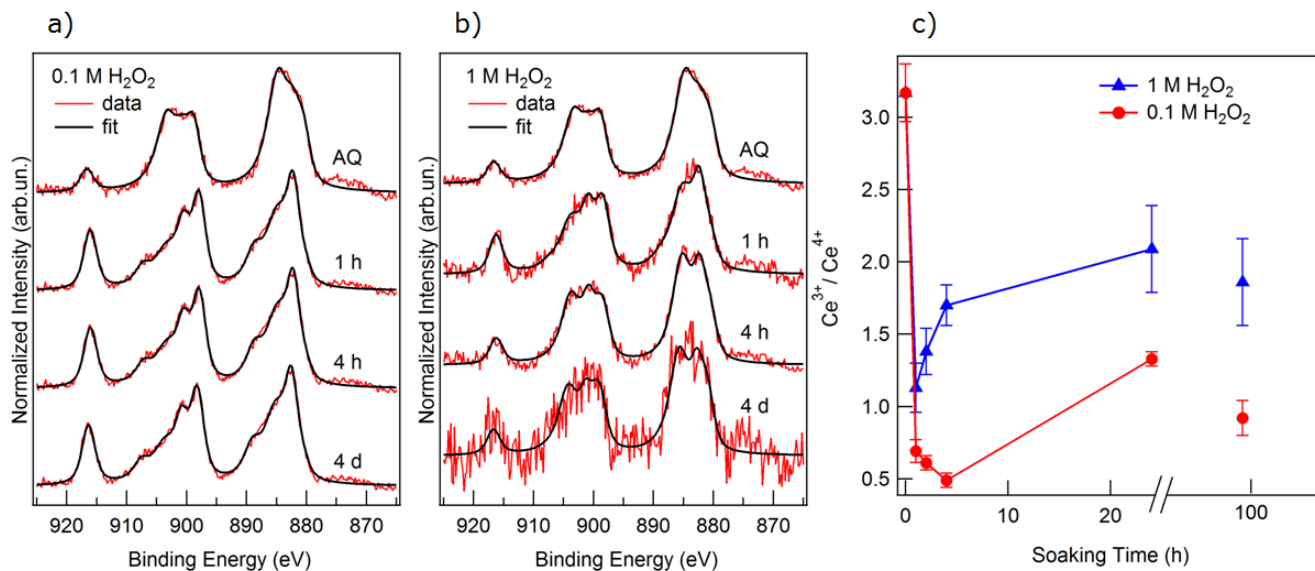
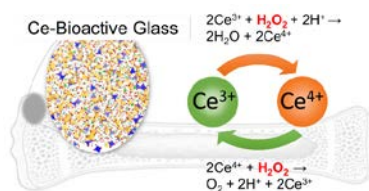


Table of Content



We report the first evidence of catalase mimetic activity of bioactive glasses doped with CeO_2 .